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NEET 2017
07-05-2017

SUBJECT - CHEMISTRY

Time : 3 :00 Hrs. समय : 3 घंटे

Max. Marks (अधिकतम अंक): 720

READ THE INSTRUCTIONS CAREFULLY (कृपया इन निर्देशों को ध्यान से पढ़ें)

Important Instructions:	महत्वपूर्ण निर्देश :
1. The Answer Sheet is inside this Test Booklet. When you are directed to open the Test Booklet, take out the Answer Sheet and fill in the particulars on Side-1 and Side-2 carefully with blue/black ball point pen only.	1. उत्तर पत्र इस परीक्षा पुस्तिका के अन्दर रखा है। जब आपको परीक्षा पुस्तिका खोलने को कहा जाए, तो उत्तर पत्र निकाल कर पृष्ठ-1 एवं पृष्ठ-2 पर केवल नीले/काले बॉल पॉइंट पेन से विवरण भरें।
2. The test is of 3 hours duration and Test Booklet contains 180 questions . Each question carries 4 marks . For each correct response, the candidate will get 4 marks . For each incorrect response, one mark will be deducted from the total scores. The maximum marks are 720 .	2. परीक्षा की अवधि 3 घंटे है एवं परीक्षा पुस्तिका में 180 प्रश्न हैं। प्रत्येक प्रश्न 4 अंक का है। प्रत्येक सही उत्तर के लिए परीक्षार्थी को 4 अंक दिए जाएंगे। प्रत्येक गलत उत्तर के लिए कुल योग में से एक अंक घटाया जाएगा। अधिकतम अंक 720 हैं।
3. Use Blue/Black Ball Point Pen only for writing particulars on this page/markings response.	3. इस पृष्ठ पर विवरण अंकित करने एवं उत्तर पत्र पर निशान लगाने के लिए केवल नीले/काले बॉल पॉइंट पेन का प्रयोग करें।
4. Rough work is to be done on the space provided for this purpose in the Test Booklet only.	4. रफ कार्य इस परीक्षा पुस्तिका में निर्धारित स्थान पर ही करें।
5. On completion of the test, the candidate must handover the Answer Sheet to the invigilator in the Room/Hall. The candidates are allowed to take away this Test Booklet with them.	5. परीक्षा सम्पन्न होने पर, परीक्षार्थी कक्ष/हॉल छोड़ने से पूर्व उत्तर पत्र कक्ष निरीक्षक को अवश्य सौंप दें। परीक्षार्थी अपने साथ प्रश्न पुस्तिका को ले जा सकते हैं।
6. The CODE for this Booklet is _____. Make sure that the CODE printed on Side-2 of the Answer Sheet is the same as that on this Booklet. In case of discrepancy, the candidate should immediately report the matter to the Invigilator for replacement of both the Test Booklets and the Answer Sheets.	6. इस पुस्तिका का संकेत है _____. यह सुनिश्चित कर लें कि इस पुस्तिका का संकेत, उत्तर पत्र के पृष्ठ-2 पद छपे संकेत से मिलता है। अगर यह भिन्न हो, तो परीक्षार्थी दूसरी परीक्षा पुस्तिका और उत्तर पत्र लेने के लिए निरीक्षक को तुरन्त अवगत कराएं।
7. The Candidates should ensure that the Answer Sheet is not folded. Do not make any stray marks on the Answer Sheet. Do not write your roll no. anywhere else except in the specified space in the Test Booklet/Answer Sheet.	7. परीक्षार्थी सुनिश्चित करें कि इस उत्तर पत्र को मोड़ा न जाए एवं उस पर कोई अन्य निशान न लगाएं। परीक्षार्थी अपना अनुक्रमांक प्रश्न पुस्तिका/उत्तर पत्र में निर्धारित स्थान के अतिरिक्त अन्यत्र न लिखें।
8. Use of white fluid for correction is NOT permissible on the Answer Sheet.	8. उत्तर पत्र पर किसी प्रकार के संशोधन हेतु व्हाइट फ्लूइड के प्रयोग की अनुमति नहीं है।

In case of any ambiguity in translation of any question, English version shall be treated as final.

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Candidate's Signature: _____ Invigilator's Signature: _____

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CHEMISTRY

136. The reason for greater range of oxidation states in actinoids is attributed to :

- (1) The radioactive nature of actinoids
- (2) Actinoid contraction
- (3) 5f, 6d and 7s levels having comparable energies
- (4) 4f and 5d levels being close in energies

Ans. (3)

137. An example of a sigma bonded organometallic compound is

- (1) Ruthenocene
- (2) Grignard's reagent
- (3) Ferrocene
- (4) Cobaltocene

Ans. (2)

138. Which one is the **wrong** statement ?

(1) de-Broglie's wavelength is given by $\lambda = \frac{h}{m\nu}$, where m = mass of the particle, ν = group velocity of the particle.

(2) The uncertainty principle is $\Delta E \times \Delta t \geq \frac{h}{4\pi}$

(3) Half filled and fully filled orbitals have greater stability due to greater exchange energy, greater symmetry and more balanced arrangement.

(4) The energy of 2s orbital is less than the energy of 2p orbital in case of Hydrogen like atoms.

Ans. (4)

139. Mixture of chloroxylenol and terpineol acts as :

- (1) Analgesic
- (2) Antiseptic
- (3) Antipyretic
- (4) Antibiotic

Ans. (2)

140. The element Z = 114 has been discovered recently. It will belong to which of the following family/group and electronic configuration ?

(1) Halogen family, $[\text{Rn}] 5f^{14}6d^{10}7s^27p^5$

(2) Carbon family, $[\text{Rn}] 5f^{14}6d^{10}7s^27p^2$

(3) Oxygen family, $[\text{Rn}] 5f^{14}6d^{10}7s^27p^4$

(4) Nitrogen family, $[\text{Rn}] 5f^{14}6d^{10}7s^27p^6$

Ans. (2)

141. A 20 litre container at 400 K contains $\text{CO}_2(\text{g})$ at pressure 0.4 atm and an excess of SrO neglect the volume of solid SrO). The volume of the container is now decreased by moving the movable piston fitted in the container. The maximum volume of the container, when pressure of CO_2 attains its maximum value, will be :

(Given that : $\text{SrCO}_3(\text{s}) \rightleftharpoons \text{SrO}(\text{s}) + \text{CO}_2(\text{g})$, $K_p = 1.6 \text{ atm}$)

(1) 5 litre

(2) 10 litre

(3) 4 litre

(4) 2 litre

Ans. (1)

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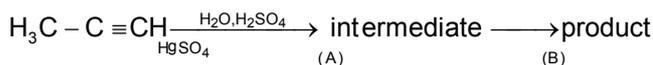
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142. Predict the correct intermediate and product in the following reaction :



- (1) A : $\text{H}_3\text{C}-\underset{\text{SO}_4}{\text{C}}=\text{CH}_2$ B : $\text{H}_3\text{C}-\underset{\text{O}}{\text{C}}-\text{CH}_3$ (2) A : $\text{H}_3\text{C}-\underset{\text{OH}}{\text{C}}=\text{CH}_2$ B : $\text{CH}_3-\underset{\text{SO}_4}{\text{C}}-\text{CH}_2$
 (3) A : $\text{CH}_3-\underset{\text{O}}{\text{C}}-\text{CH}_3$ B : $\text{H}_3\text{C}-\text{C}\equiv\text{CH}$ (4) A : $\text{H}_3\text{C}-\underset{\text{OH}}{\text{C}}=\text{CH}_2$ B : $\text{H}_3\text{C}-\underset{\text{O}}{\text{C}}-\text{CH}_3$

Ans. (4)

143. Which of the following is a sink for CO ?

- (1) Haemoglobin (2) Micro organisms present in the soil.
 (3) Oceans (4) Plants

Ans. (1)

144. Which of the following reactions is appropriate for converting acetamide to methanamine ?

- (1) Carbylamine reaction (2) Hoffmann hypobromamide reaction
 (3) Stephens reaction (4) Gabriels phthalimide synthesis

Ans. (2)

145. The species, having bond angles of 120° is :

- (1) PH_3 (2) ClF_3 (3) NCl_3 (4) BCl_3

Ans. (4)

146. The correct of order the stoichiometries of AgCl formed when AgNO_3 in excess is treated with the complexes $\text{CoCl}_3 \cdot 6\text{NH}_3$, $\text{CoCl}_3 \cdot 5\text{NH}_3$, $\text{CoCl}_3 \cdot 4\text{NH}_3$ respectively is :

- (1) 1 AgCl , 3 AgCl , 2 AgCl (2) 3 AgCl , 1 AgCl , 2 AgCl
 (3) 3 AgCl , 2 AgCl , 1 AgCl (4) 2 AgCl , 3 AgCl , 1 AgCl

Ans. (3)

147. For a given reaction $\Delta H = 35.5 \text{ kJ mol}^{-1}$ and $\Delta S = 83.6 \text{ JK}^{-1} \text{ mol}^{-1}$. The reaction is spontaneous at : (Assume that ΔH and ΔS do not vary with temperature)

- (1) $T < 425 \text{ K}$ (2) $T > 425 \text{ K}$ (3) All temperatures (4) $T > 298 \text{ K}$

Ans. (2)

148. Match the interhalogen compounds of Column I with the geometry in column II and Assign the correct code.

- | Column I | Column II |
|--------------------|-----------------------------|
| (a) XX' | (i) T-shape |
| (b) XX_3' | (ii) Pentagonal bipyramidal |
| (c) XX_5' | (iii) Linear |
| (d) XX_7' | (iv) Square-pyramidal |
| | (v) Tetrahedral |

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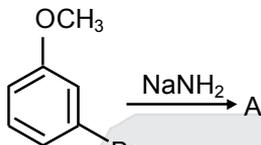
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Code :

	(a)	(b)	(c)	(d)		(a)	(b)	(c)	(d)
(1)	(iii)	(iv)	(i)	(ii)	(2)	(iii)	(i)	(iv)	(ii)
(3)	(v)	(iv)	(iii)	(ii)	(4)	(iv)	(iii)	(ii)	(i)

Ans. (2)

149. Identify A and predict the type of reaction :



- (1) and substitution reaction
- (2) and elimination addition reaction
- (3) and cine substitution reaction
- (4) and cine substitution reaction

Ans. (2)

150. Which one of the following statements is not correct?

- (1) Catalyst does not initiate any reaction.
 (2) The value of equilibrium constant is changed in the presence of a catalyst in the reaction at equilibrium.
 (3) Enzymes catalyse mainly bio-chemical reactions
 (4) Coenzymes increase the catalytic activity of enzyme.

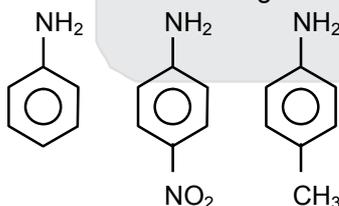
Ans. (2)

151. Name the gas that can readily decolourise acidified KMnO_4 solution:

- (1) CO_2 (2) SO_2 (3) NO_2 (4) P_2O_5

Ans. (2)

152. The correct increasing order of basic strength for the following compounds is :



- (I) (II) (III)
- (1) II < III < I (2) III < I < II (3) III < II < I (4) II < I < III

Ans. (4)

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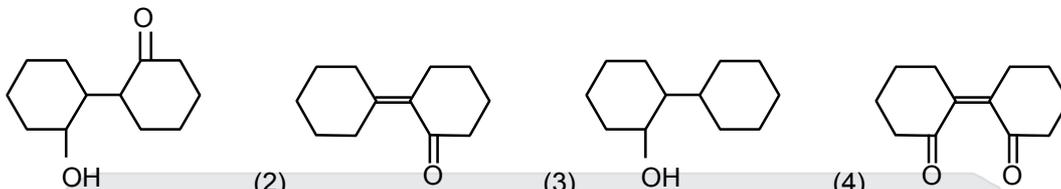
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153. If molality of the dilute solution is doubled, the value of molal depression constant (K_f) will be :
(1) doubled (2) halved (3) tripled (4) unchanged

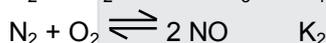
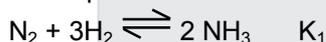
Ans. (4)

154. Of the following, which is the product formed when cyclohexanone undergoes aldol condensation followed by heating?

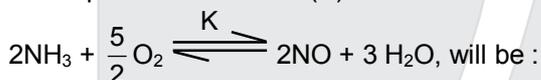


Ans. (2)

155. The equilibrium constants of the following are :



The equilibrium constant (K) of the reaction :



(1) $K_1 K_3^3 / K_2$

(2) $K_2 K_3^3 / K_1$

(3) $K_2 K_3 / K_1$

(4) $K_2^3 K_3 / K_1$

Ans. (2)

156. The correct statement regarding electrophile is :

- (1) Electrophile is a negatively charged species and can form a bond by accepting a pair of electrons from a nucleophile
- (2) Electrophile is a negatively charged species and can form a bond by accepting a pair of electrons from another electrophile
- (3) Electrophiles are generally neutral species and can form a bond by accepting a pair of electrons from a nucleophile
- (4) Electrophile can be either neutral or positively charged species and can form a bond by accepting a pair of electrons from a nucleophile

Ans. (4)

157. A gas is allowed to expand in a well insulated container against a constant external pressure of 2.5 atm from an initial volume of 2.50 L to a final volume of 4.50 L. The change in internal energy ΔU of the gas in joules will be:

- (1) 1136.25 J (2) -500 J (3) -505 J (4) +505 J

Ans. (3)

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158. Which of the following pairs of compounds is isoelectronic and isostructural?

- (1) BeCl_2 , XeF_2 (2) TeI_2 , XeF_2 (3) IBr_2^- , XeF_2 (4) IF_3 , XeF_2

Ans. (3)

159. Which of the incorrect statement?

- (1) $\text{FeO}_{0.98}$ has non stoichiometric metal deficiency defect.
 (2) Density decreases in case of crystals with Schottky's defect.
 (3) NaCl(s) is insulator, silicon is semiconductor, silver is conductor, quartz is piezo electric crystal.
 (4) Frenkel defect is favoured in those ionic compounds in which sizes of cation and anions are almost equal.

Ans. (1, 4)

160. The heating of phenyl-methyl ethers with HI produces.

- (1) ethyl chlorides (2) iodobenzene (3) phenol (4) benzene

Ans. (3)

161. Correct increasing order for the wavelength of absorption in the visible region for the complexes of Co^{3+} is :

- (1) $[\text{Co(en)}_3]^{3+}$, $[\text{Co(NH}_3)_6]^{3+}$, $[\text{Co(H}_2\text{O)}_6]^{3+}$ (2) $[\text{Co(H}_2\text{O)}_6]^{3+}$, $[\text{Co(en)}_3]^{3+}$, $[\text{Co(NH}_3)_6]^{3+}$
 (3) $[\text{Co(H}_2\text{O)}_6]^{3+}$, $[\text{Co(NH}_3)_6]^{3+}$, $[\text{Co(en)}_3]^{3+}$ (4) $[\text{Co(NH}_3)_6]^{3+}$, $[\text{Co(en)}_3]^{3+}$, $[\text{Co(H}_2\text{O)}_6]^{3+}$

Ans. (1)

162. Pick out the correct statement with respect to $[\text{Mn(CN)}_6]^{3-}$:

- (1) It is sp^3d^2 hybridised and octahedral (2) It is sp^3d^2 hybridised an tetrahedral
 (3) It is d^2sp^3 hybridised and octahedral (4) It is dsp^2 hybridised and square planar.

Ans. (3)

163. With respect to the conformers of ethane, which of the following statements is true?

- (1) Bond angle remains same but bond length changes
 (2) Bond angle changes but bond length remains same
 (3) Both bond angle and bond length changes
 (4) Both bond angles and bond length remains same

Ans. (4)

164. Which of the following is dependent on temperature ?

- (1) Molality (2) Molarity (3) Mole fraction (4) Weight percentage

Ans. (2)

165. Which of the following statement is not correct ?

- (1) Insulin maintains sugar level in the blood of a human body.
 (2) Ovalbumin is a simple food reserve in egg white
 (3) Blood proteins thrombin and fibrinogen are involved in blood clotting.
 (4) Denaturation makes the proteins more active.

Ans. (4)

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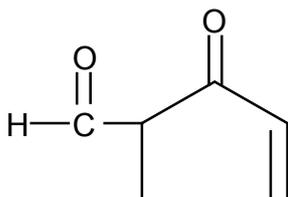
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166. The IUPAC name of the compound is



- (1) 3-keto-2-methylhex-4-enal (2) 5-formylhex-2-en-3-one
(3) 5-methyl-4-oxohex-2-en-5-al (4) 3-keto-2-methylhex-5-enal

Ans. (1)

167. HgCl_2 and I_2 both when dissolved in water containing I^- ions the pair of species formed is :

- (1) $\text{HgI}_2, \text{I}_3^-$ (2) HgI_2, I^- (3) $\text{HgI}_4^{2-}, \text{I}_3^-$ (4) $\text{Hg}_2\text{I}_2, \text{I}^-$

Ans. (3)

168. It is because of inability of ns^2 electrons of the valence shell to participate in bonding that :

- (1) Sn^{2+} is reducing while Pb^{4+} is oxidising
(2) Sn^{2+} is oxidising while Pb^{4+} is reducing
(3) Sn^{2+} and Pb^{2+} are both oxidising and reducing
(4) Sn^{4+} is reducing while Pb^{4+} is oxidising

Ans. (1)

169. Mechanism of a hypothetical reaction $\text{X}_2 + \text{Y}_2 \rightarrow 2\text{XY}$ is given below :

- (i) $\text{X}_2 \rightarrow \text{X} + \text{X}$ (fast)
(ii) $\text{X} + \text{Y}_2 \rightleftharpoons \text{XY}$ (Slow)
(iii) $\text{X} + \text{Y} \rightarrow \text{XY}$ (fast)

The overall order of the reaction will be

- (1) 1 (2) 2 (3) 0 (4) 1.5

Ans. (1,4)

170. Concentration of the Ag^+ ions in a saturated solution of $\text{Ag}_2\text{C}_2\text{O}_4$ is $2.2 \times 10^{-4} \text{ molL}^{-1}$. Solubility product of $\text{Ag}_2\text{C}_2\text{O}_4$ is :

- (1) 2.42×10^{-8} (2) 2.66×10^{-12} (3) 4.5×10^{-11} (4) 5.3×10^{-12}

Ans. (4)

171. Extraction of gold and silver involves leaching with CN^- ion. Silver is later recovered by :

- (1) liquation (2) distillation (3) zone refining (4) displacement with Zn

Ans. (4)

172. Which one is the correct order of acidity ?

- (1) $\text{CH}_2=\text{CH}_2 > \text{CH}_3-\text{CH}=\text{CH}_2 > \text{CH}_3-\text{C}\equiv\text{CH} > \text{CH}\equiv\text{CH}$
(2) $\text{CH}\equiv\text{CH} > \text{CH}_3-\text{C}\equiv\text{CH} > \text{CH}_2=\text{CH}_2 > \text{CH}_3-\text{CH}_3$
(3) $\text{CH}\equiv\text{CH} > \text{CH}_2=\text{CH}_2 > \text{CH}_3-\text{C}\equiv\text{CH} > \text{CH}_3-\text{CH}_3$
(4) $\text{CH}_3-\text{CH}_3 > \text{CH}_2=\text{CH}_2 > \text{CH}_3-\text{C}\equiv\text{CH} > \text{CH}\equiv\text{CH}$

Ans. (2)

173. Ionic mobility of which of the following alkali metal ions is lowest when aqueous solution of their salts are put under an electric field ?

- (1) Na (2) K (3) Rb (4) Li

Ans. (4)

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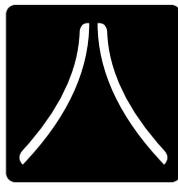
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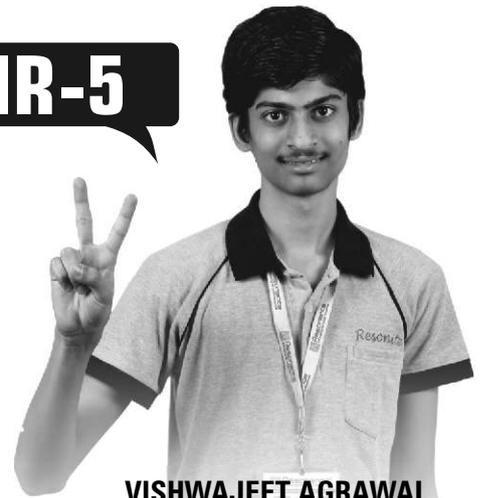
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