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JEE (Main)

PAPER-1 (B.E./B. TECH.)

2023

COMPUTER BASED TEST (CBT) Questions & Solutions

Date: 24 January, 2023 (SHIFT-1) | TIME : (9.00 a.m. to 12.00 p.m.)

Duration: 3 Hours | Max. Marks: 300

SUBJECT: CHEMISTRY

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PART : CHEMISTRY

31. In the depression of freezing point experiment
A. Vapour pressure of the solution is less than that of pure solvent
B. Vapour pressure of the solution is more than that of pure solvent
C. Only solute molecules solidify at the freezing point
D. Only solvent molecules solidify at the freezing point

Choose the most appropriate answer from the options given below:

- (1) A only (2) B and C only (3) A and C only (4) A and D only

NTA. (4)

RESO. (4)

Sol. In the depression of freezing point experiment

- A. Vapour pressure of the solution is less than that of pure solvent
B. Only solvent molecule get solidify.

32. Order of Covalent bond;

- A. $KF > KI$; $LiF > KF$
B. $KF < KI$; $LiF > KF$
C. $SnCl_4 > SnCl_2$; $CuCl > NaCl$
D. $LiF > KF$; $CuCl < NaCl$
E. $KF < KI$; $CuCl > NaCl$

Choose the correct answer from the options given below:

- (1) A, B only (2) B, C only (3) B, C, E only (4) C, E only

NTA. (3)

RESO. (3)

Sol. Order of covalent character bond is

- B. $KF < KI$ & $LiF > KF$
C. $SnCl_4 > SnCl_2$ & $CuCl > NaCl$
E. $KF < KI$ & $CuCl > NaCl$

33. Match List I with List II.

LIST- I		LIST-II	
A	Reverberatory furnace	I	Pig Iron
B	Electrolytic cell	II	Aluminum
C	Blast furnace	III	Silicon
D	Zone Refining furnace	IV	Copper

Choose the correct answer from the options given below:

- (1) A-I, B-III, C-II, D-IV (2) A-III, B-IV, C-I, D-II (3) A-I, B-IV, C-II, D-III (4) A-IV, B-II, C-I, D-III

NTA. (4)

RESO. (4)

34. Which of the following is true about freons?

- (1) These are radicals of chlorine and chlorine monoxide
(2) These are chlorofluorocarbon compounds
(3) All radicals are called freons
(4) These are chemicals causing skin cancer

NTA. (2)

RESO. (2)

Sol. Freons are chlorofluoro carbons

Ex. Freon 012, 112 etc.

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35. Match List I with List II

LIST-I		LIST-II	
A	Chlorophyll	I	Na ₂ CO ₃
B	Soda ash	II	CaSO ₄
C	Dentistry, Ornamental work	III	Mg ²⁺
D	Used in white washing	IV	Ca(OH) ₂

Choose the correct answer from the options given below:

- (1) A-III, B-I, C-II, D-IV (2) A-II, B-III, C-IV, D-I
(3) A-III, B-IV, C-I, D-II (4) A-II, B-I, C-III, D-IV

NTA. (1)

RESO. (1)

Sol. Theory Based

36. The magnetic moment of a transition metal compound has been calculated to be 3.87 B.M. The metal ion is

- (1) Ti²⁺ (2) Mn²⁺ (3) V²⁺ (4) Cr²⁺

NTA. (3)

RESO. (3)

Sol.

Ion	Outer electronic configuration	no. of unpaired electron	Magnetic moment
Mn ²⁺	3d ⁵ 4s ⁰	5	$\sqrt{35}$ BM
V ²⁺	3d ³ 4s ⁰	3	$\sqrt{15}$ BM
Ti ²⁺	3d ² 4s ⁰	2	$\sqrt{8}$ BM
Cr ²⁺	3d ⁴ 4s ⁰	4	$\sqrt{24}$ BM

37. Decreasing order of the hydrogen bonding in following forms of water is correctly represented by

- A. Liquid water B. Ice C. Impure water

Choose the correct answer from the options given below:

- (1) C > B > A (2) B > A > C (3) A = B > C (4) A > B > C

NTA. (2)

RESO. (2)

Sol. In ice each H₂O molecule forms four H-bonds & is tetrahedrally structured so strength of H bond is greater in ice than in pure water.

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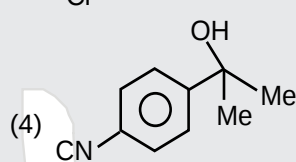
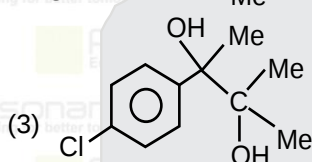
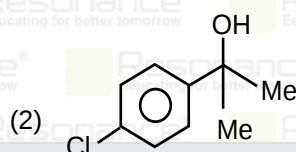
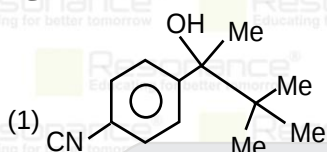
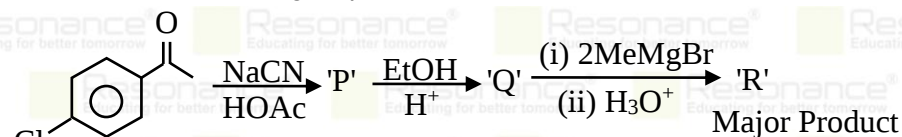
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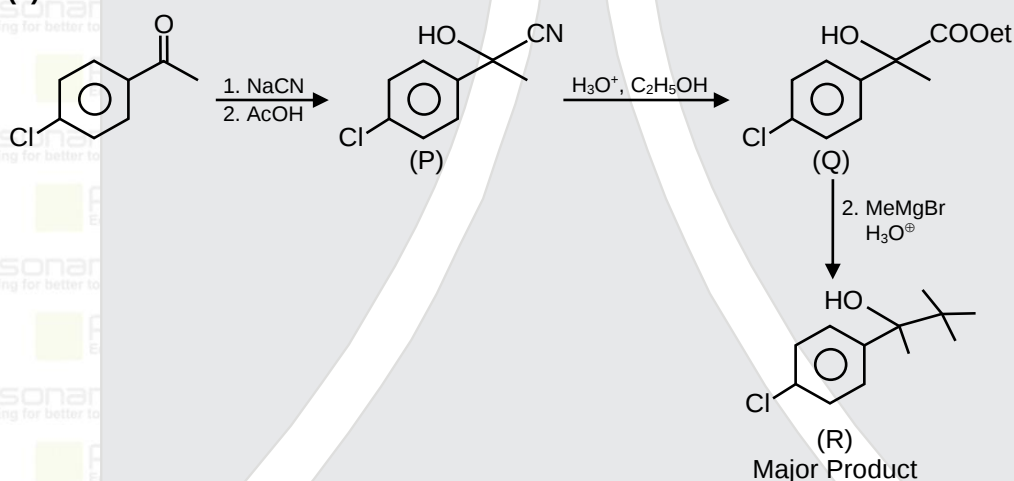
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38. 'R' formed in the following sequence of reactions is :

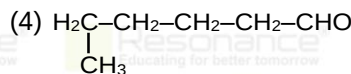
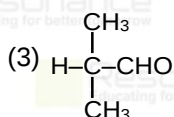
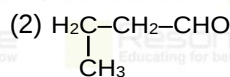
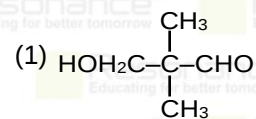
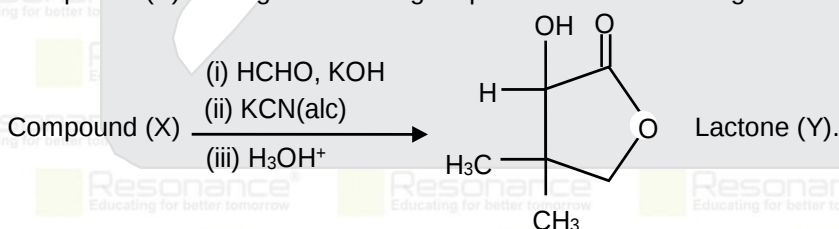


NTA. (3)
RESO. (3)

Sol.



39. Compound (X) undergoes following sequence of reactions to give the Lactone (Y).



NTA. (3)
RESO. (3)

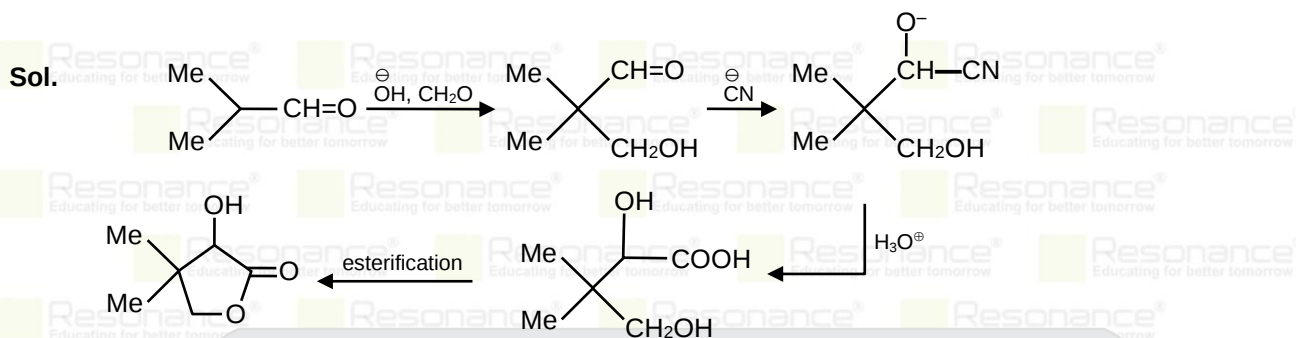
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40. An ammoniacal metal salt solution gives a brilliant red precipitate on addition of dimethylglyoxime. The metal ion is:

(1) Co^{2+} (2) Fe^{2+} (3) Ni^{2+} (4) Cu^{2+}

NTA. (3)

RESO. (3)

Sol. $\text{Ni}^{2+} + \text{dmg} + \text{NH}_4\text{OH} \rightarrow [\text{Ni}(\text{dmg})_2] \downarrow + \text{NH}_4^+ + \text{H}_2\text{O}$
red ppt

41. Reaction of BeO with ammonia and hydrogen fluoride gives A which on thermal decomposition gives BeF_2 and NH_4F . What is 'A' ?

(1) $(\text{NH}_4)\text{BeF}_3$ (2) H_3NBeF_3 (3) $(\text{NH}_4)\text{Be}_2\text{F}_5$ (4) $(\text{NH}_4)_2\text{BeF}_4$

NTA. (4)

RESO. (4)

Sol. Thermal decomposition of $(\text{NH}_4)_2\text{BeF}_4$ is the best route for preparation of BeF_2

42. **Assertion A :** Hydrolysis of an alkyl chloride is a slow reaction but in the presence of NaI , the rate of the hydrolysis increases.

Reason R : I^- is a good nucleophile as well as a good leaving group.

In the light of the above statements, choose the correct answer from the options given below

- (1) A is false but R is true
(2) A is true but R is false
(3) Both A and R are true but R is NOT the correct explanation of A
(4) Both A and R are true and R is the correct explanation of A

NTA. (4)

RESO. (4)

Sol. I^- being good nucleophile form R-I which get hydrolysed readily

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43. Given below are two statements:

Statement I : Noradrenaline is a neurotransmitter.

Statement II : Low level of noradrenaline is not the cause of depression in human. In the light of the above statements, choose the correct answer from the options given below

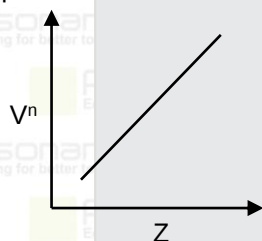
- (1) Statement I is incorrect but Statement II is correct
- (2) Both Statement I and Statement II are correct
- (3) Statement I is correct but Statement II is incorrect
- (4) Both Statement I and Statement II are incorrect

NTA. (3)

RESO. (3)

Sol. Low level of noradrenaline causes depression
(NCERT Part-II, XII class)

44. It is observed that characteristic X-ray spectra of elements show regularity. When frequency to the power "n" i.e. ν^n of X-rays emitted is plotted against atomic number "Z", following graph is obtained.



The value of "n" is

- (1) 2
- (2) 3
- (3) 1
- (4) $\frac{1}{2}$

NTA. (4)

RESO. (4)

Sol.

$$\nu^n \propto Z$$

$$\sqrt{\nu} = a(Z - b)$$

$$\therefore n = \frac{1}{2}$$

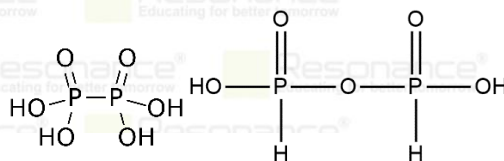
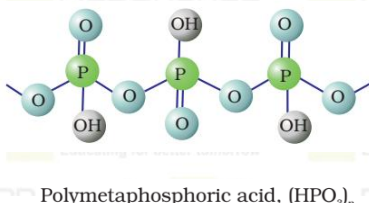
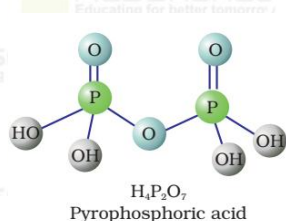
45. Which of the Phosphorus oxoacid can create silver mirror from AgNO_3 solution?

- (1) $\text{H}_4\text{P}_2\text{O}_7$
- (2) $(\text{HPO}_3)_n$
- (3) $\text{H}_4\text{P}_2\text{O}_6$
- (4) $\text{H}_4\text{P}_2\text{O}_5$

NTA. (4)

RESO. (4)

Sol. The acids which contain P-H bond have strong reducing properties. Thus, $\text{H}_4\text{P}_2\text{O}_5$ acid is act as good reducing agent as it contains P-H bonds and reduces AgNO_3 to Ag.



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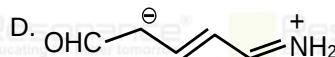
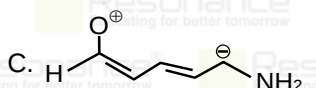
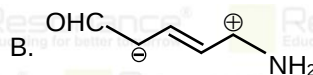
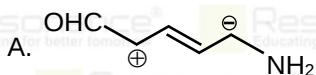
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46. Increasing order of stability of the resonance structures is:



Choose the correct answer from the options given below:

(1) D, C, A, B

(2) C, D, B, A

(3) D, C, B, A

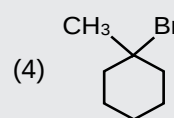
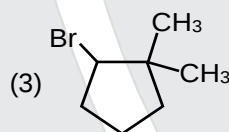
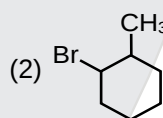
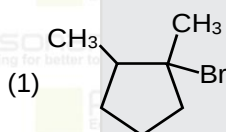
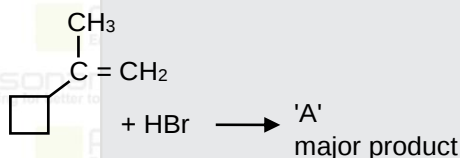
(4) C, D, A, B

NTA. (4)

RESO. (None of the given options are correct)

Correct answer should be $C < A < B < D$

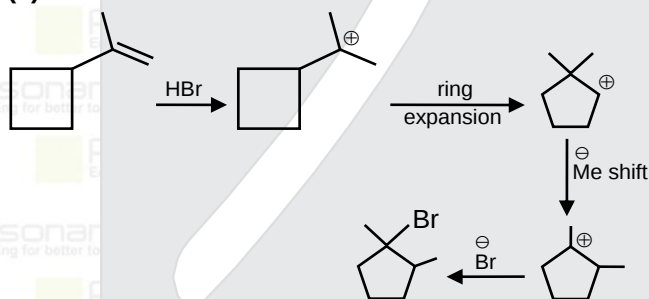
47. In the following given reaction 'A' is



NTA. (1)

RESO. (1)

Sol.



48. The primary and secondary valencies of cobalt respectively in $[\text{Co}(\text{NH}_3)_5\text{Cl}]\text{Cl}_2$ are:

(1) 3 and 5

(2) 2 and 6

(3) 2 and 8

(4) 3 and 6

NTA. (4)

RESO. (4)

Sol. Primary valency = Oxidation number of central metal.

Secondary valency = Co-ordination number of central metal

In $[\text{Co}(\text{NH}_3)_5\text{Cl}]\text{Cl}_2$

Primary valency = 3

Secondary valency = 6

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49. Statement I: For colloidal particles, the values of colligative properties are of small order as compared to values shown by true solutions at same concentration.

Statement II: For colloidal particles, the potential difference between the fixed layer and the diffused layer of same charges is called the electrokinetic potential or zeta potential. In the light of the above statements, choose the correct answer from the options given below

- (1) Statement I is false but Statement II is true
- (2) Both Statement I and Statement II are true
- (3) Statement I is true but Statement II is false
- (4) Both Statement I and Statement II are false

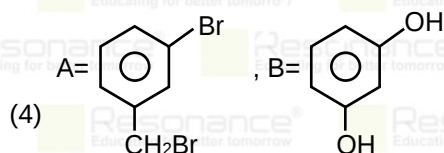
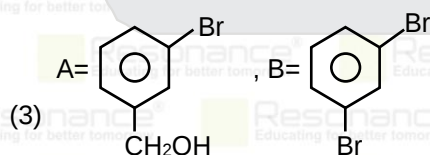
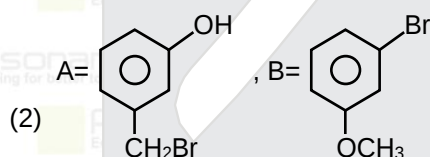
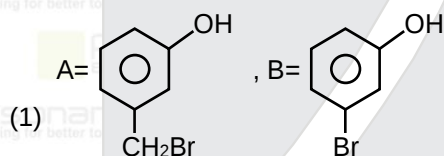
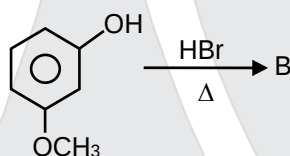
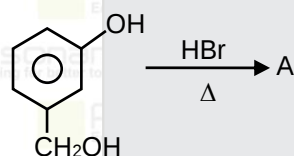
NTA. (3)

RESO. (3)

Sol. **Statement I :** For colloidal particles, the values of colligative properties are small order as compared to values shown by true solutions at same concentration.

Statement II : The potential difference between the fixed layer and diffused layer of opposite charge is called electro kinetics potential or zeta potential.

50. 'A' and 'B' formed in the following set of reactions are:



NTA. (4)

RESO. (4)

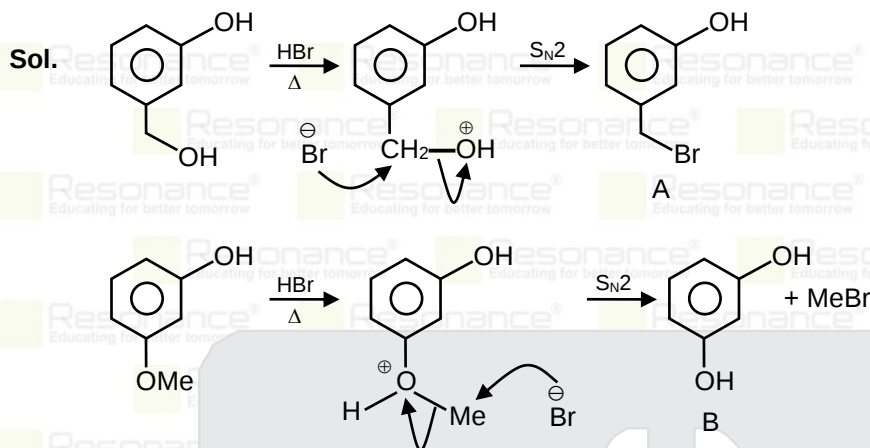
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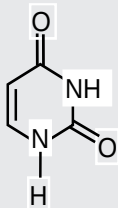
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51. Uracil is a base present in RNA with the following structure. % of N in uracil is



Given:

Molar mass N = 14 g mol⁻¹

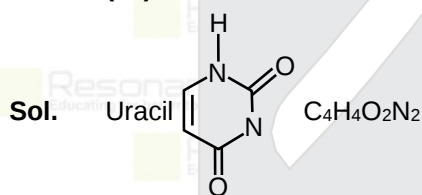
O = 16 g mol⁻¹

C = 12 g mol⁻¹

H = 1 g mol⁻¹

NTA. (25)

RESO. (25)



$$\frac{2 \times 14}{4 \times 14 + 4 \times 1 + 2 \times 16 + 2 \times 14} \times 100 = 25\%$$

52. If wavelength of the first line of the Paschen series of hydrogen atom is 720 nm, then the wavelength of the second line of this series is _____ nm. (Nearest integer)

NTA. (492)

RESO. (492)

Sol.

$$\frac{1}{\lambda} = R_H \times \left[\frac{1}{n_1^2} - \frac{1}{n_2^2} \right] \times (Z)^2$$

For 1st line of Paschen series : $n_1 = 3$; $n_2 = 4$

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$$\frac{1}{720} = \frac{1}{\lambda_1} = R_H \times (Z)^2 \times \left[\frac{1}{(3)^2} - \frac{1}{(4)^2} \right] \quad \dots(1)$$

$$\frac{1}{720} = \frac{5}{9 \times 16} \times R_H \times (Z)^2$$

For IInd line of Paschen series : $n_1 = 3$; $n_2 = 5$

$$\frac{1}{\lambda_2} = R_H \times (Z)^2 \times \left[\frac{1}{(3)^2} - \frac{1}{(5)^2} \right] \quad \dots(2)$$

Dividing (1) by (2)

$$\frac{\lambda_2}{720} = \frac{7}{9 \times 16} \times \frac{9 \times 25}{16}$$

$$\lambda_2 = 492 \text{ nm}$$

53. At 298 K, a 1 litre solution containing 10 mmol of $\text{Cr}_2\text{O}_7^{2-}$ and 100 mmol of Cr^{3+} shows a pH of 3.0.

Given: $\text{Cr}_2\text{O}_7^{2-} \rightarrow \text{Cr}^{3+}$; $E^\circ = 1.33\text{V}$

$$\text{and } \frac{(2.303RT)}{F} = 0.059\text{V}$$

The potential for the half cell reaction is $x \times 10^{-3}\text{V}$. The value of x is

NTA. (917)

RESO. (917)

54. The dissociation constant of acetic acid is $x \times 10^{-5}$. When 25 mL of 0.2 M CH_3COONa solution is mixed with 25 mL of 0.02 M CH_3COOH solution, the pH of the resultant solution is found to be equal to 5. The value of x is

NTA. (10)

RESO. (10)

55. For independent processes at 300 K

Process	$\Delta H/\text{kJ mol}^{-1}$	$\Delta S/\text{J mol}^{-1}$
A	-25	-80
B	-22	40
C	25	-50
D	22	20

The number of non-spontaneous processes from the following is

NTA. (2)

RESO. (2)

Sol.

Process	sign of ΔG	Spontaneity of process
A	(-) ve	Spontaneous
B	(-)	Spontaneous
C	(+)	Non spontaneous
D	(+)	non spontaneous

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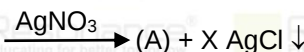
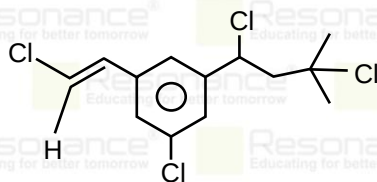
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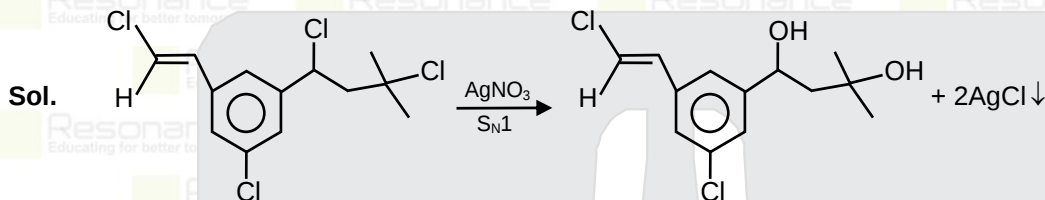
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56. Number of moles of AgCl formed in the following reaction is



NTA. (2)

RESO. (2)



57. The d-electronic configuration of $[\text{CoCl}_4]^{2-}$ in tetrahedral crystal field is $e^m t_2^n$. Sum of "m" and "number of unpaired electrons" is _____

NTA. (7)

RESO. (7)

58. 5 g of NaOH was dissolved in deionized water to prepare a 450 mL stock solution. What volume (in mL) of this solution would be required to prepare 500 mL of 0.1 M solution?
Given: Molar Mass of Na, O and H is 23, 16 and 1 g mol⁻¹ respectively

NTA. (180)

RESO. (180)

Sol. $(\text{Molarity})_i = \left\{ \frac{5 \times 1000}{40 \times 450} \right\}$

By dilution law

$$M_i V_i = M_f V_f$$

$$\Rightarrow \left\{ \frac{5 \times 1000}{40 \times 450} \right\} \times V = 0.1 \times 500$$

$$V = \frac{50 \times 40 \times 450}{5 \times 1000} = 5 \times 4 \times 9 = 180 \text{ mL}$$

59. The number of correct statement/s from the following is _____

- A. Larger the activation energy, smaller is the value of the rate constant.
B. The higher is the activation energy, higher is the value of the temperature coefficient.
C. At lower temperatures, increase in temperature causes more change in the value of k than at higher temperature
D. A plot of $\ln k$ vs $\frac{1}{T}$ is a straight line with slope equal to $-\frac{E_a}{R}$

NTA. (3)

RESO. (3)

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- Sol.** A. Larger the activation energy, smaller is the value of the rate constant $\Rightarrow$ true statement
B. The higher is the activation energy, higher is the value of the temperature coefficient $\Rightarrow$ true statement

$$D. \ln K = \ln A - \left(\frac{E_a}{R} \right) T \quad \text{Slope} = - \left(\frac{E_a}{R} \right)$$

60. When $\text{Fe}_{0.93}\text{O}$ is heated in presence of oxygen, it converts to Fe_2O_3 . The number of correct statement/s from the following is _____

- A. The equivalent weight of $\text{Fe}_{0.93}\text{O}$ is $\frac{\text{Molecular weight}}{0.79}$
B. The number of moles of Fe^{2+} and Fe^{3+} in 1 mole of $\text{Fe}_{0.93}\text{O}$ is 0.79 and 0.14 respectively
C. $\text{Fe}_{0.93}\text{O}$ is metal deficient with lattice comprising of cubic closed packed arrangement of O^{2-} ions
D. The % composition of Fe^{2+} and Fe^{3+} in $\text{Fe}_{0.93}\text{O}$ is 85% and 15% respectively

NTA. (4)

RESO. (4)

Sol. 1 mole $\text{Fe}_{0.93}\text{O}$ contain 0.79 mole Fe^{2+} and 0.14 mole Fe^{3+}

% of Fe^{3+} = 15% & % of Fe^{2+} = 85%

$$\text{Equivalent weight of } \text{Fe}_{0.93}\text{O} = \left\{ \frac{\text{Molecular weight}}{0.79} \right\}$$

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