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JEE

(Main)

PAPER-1 (B.E./B. TECH.)

2023

COMPUTER BASED TEST (CBT) Questions & Solutions

Date: 06 April, 2023 (SHIFT-1) | TIME : (9.00 a.m. to 12.00 p.m)

Duration: 3 Hours | Max. Marks: 300

SUBJECT: CHEMISTRY

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PART : CHEMISTRY

61. Match List I with List II :

	List I		List II
(A)	Hell-Volhard-Zelinsky reaction	(I)	NaOH + I ₂
(B)	Iodoform reaction	(II)	(i) CrO ₂ Cl ₂ , CS ₂ (ii) H ₂ O
(C)	Etard reaction	(III)	(i) Br ₂ /red phosphorus (ii) H ₂ O
(D)	Gatterman-Koch reaction	(IV)	CO, HCl, anhyd. AlCl ₃

Choose the correct answer from the options given below :

(1) A – III, B – I, C – II, D – IV

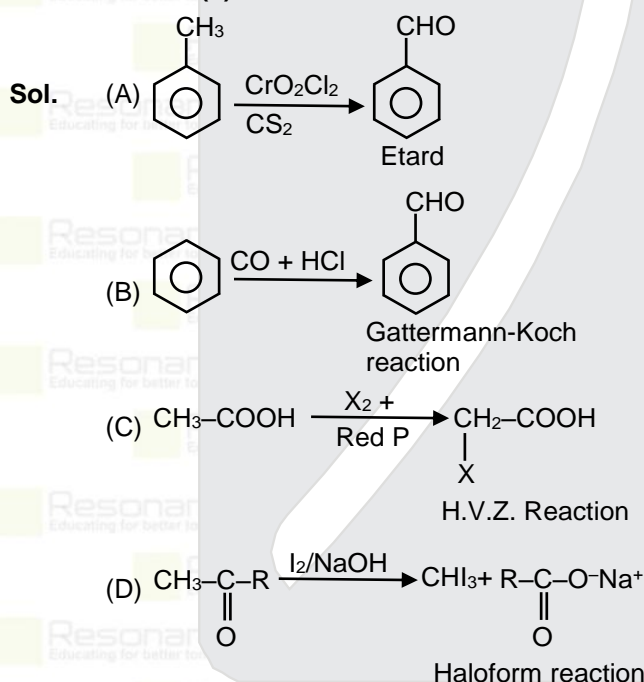
(1) A – I, B – II, C – III, D – IV

(3) A – III, B – I, C – IV, D – II

(4) A – III, B – II, C – I, D – IV

Ans. NTA - (1)

Reso - (1)



62. Given below are two statements, one is labelled as Assertion A and the other is labelled as Reason R.

Assertion A : The spin only magnetic moment value for [Fe(CN)₆]³⁻ is 1.74 BM, whereas for [Fe(H₂O)₆]³⁺ is 5.92 BM.

Reason R : In both complexes, Fe is present in +3 oxidation state.

In the light of the above statements, choose the correct answer from the options given below :

- (1) Both A and R are true and R is the correct explanation of A
- (2) Both A and R are true but R is NOT the correct explanation of A
- (3) A is true but R is false
- (4) A is false but R is true

Ans. NTA - (2)

Reso - (2)

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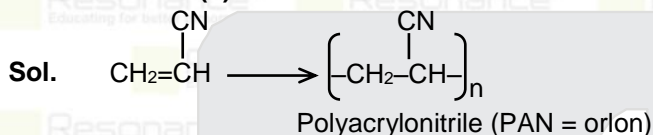
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Sol. In $[\text{Fe}(\text{H}_2\text{O})_6]^{+3}$, H_2O is a weak field ligand so configuration of Fe^{+3} is $t_{2g}^{1,1,1}, e_g^{0,0}$
so number of unpaired electron are 5 and magnetic momentum is 5.92 BM
In $[\text{Fe}(\text{CN})_6]^{-3}$, CN^- is a strong field ligand so configuration of Fe^{+3} is $t_{2g}^{2,2,1}, e_g^{0,0}$
so number of unpaired electron are 1 and magnetic momentum is 1.73 BM

63. Polymer used in orlon is :
(1) Polycarbonate (2) Polyacrylonitrile (3) Polyethene (4) Polyamide

Ans. NTA - (2)
Reso - (2)



64. The setting time of Cement is increased by adding
(1) Gypsum (2) Silica (3) Clay (4) Limestone

Ans. NTA - (1)
Reso - (1)

Sol. The purpose of adding gypsum is only to slow down the process of setting of the cement so that it gets sufficiently hardened.

65. Match List I with List II :

	List I (Vitamin)		List II (Deficiency disease)
(A)	Vitamin A	(I)	Beri-Beri
(B)	Thiamine	(II)	Cheilosis
(C)	Ascorbic acid	(III)	Xerophthalmia
(D)	Riboflavin	(IV)	Scurvy

Choose the correct answer from the options given below :

- (1) A - IV, B - II, C - III, D - I (1) A - III, B - II, C - IV, D - I
(3) A - IV, B - I, C - III, D - II (4) A - III, B - I, C - IV, D - II

Ans. NTA - (4)
Reso - (4)

Sol. Vitamin A → Xerophthalmia
Thiamine → Beri-Beri
Ascorbic acid → Scurvy
Riboflavin → Cheilosis

66. Match List I with List II :

	List I (Element detected)		List II (Reagent used/Product formed)
(A)	Nitrogen	(I)	$\text{Na}_2[\text{Fe}(\text{CN})_5\text{NO}]$
(B)	Sulphur	(II)	AgNO_3
(C)	Phosphorous	(III)	$\text{Fe}_4[\text{Fe}(\text{CN})_6]_3$
(D)	Halogen	(IV)	$(\text{NH}_4)_2\text{MoO}_4$

Choose the correct answer from the options given below :

- (1) A - IV, B - II, C - I, D - III (1) A - II, B - IV, C - I, D - III
(3) A - II, B - I, C - IV, D - III (4) A - III, B - I, C - IV, D - II

Ans. NTA - (4)
Reso - (4)

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- Sol. (1) Halogen + AgNO₃ → AgX ppt
 (2) 'S' + Sodium nitro prusside → Violet colour
 (3) 'N' + [Fe(CN)₆]⁴⁻ → Blue colour
 (4) 'P' + (NH₄)₂ MoO₄ → (NH₄)₃ PO₄.12MoO₃ → Yellow colour

67. A compound is formed by two elements X and Y. The element Y forms cubic close packed arrangement and those of element X occupy one third of the tetrahedral voids. What is the formula of the compound ?

- (1) X₂Y₃ (2) X₃Y₂ (3) X₃Y (4) XY₃

Ans. NTA - (1)

Reso - (1)

Sol. Effective number of Y atom = 4 (Y forms cubic close packed arrangement)

$$\text{Effective number of X atom} = \frac{1}{3} \times 8 = \frac{8}{3} \quad (\text{X occupy one third of the tetrahedral voids})$$

$$\text{so formula of compound} = \left(X_{\frac{8}{3}} Y_4 \right) \times \frac{3}{4} = X_2 Y_3$$

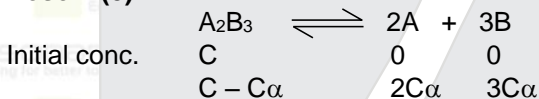
68. For a concentrated solution of a weak electrolyte (K_{eq} = equilibrium constant) A₂B₃ of concentration 'c', the degree of dissociation 'α' is

- (1) $\left(\frac{K_{eq}}{6c^5} \right)^{1/5}$ (2) $\left(\frac{K_{eq}}{5c^4} \right)^{1/5}$ (3) $\left(\frac{K_{eq}}{108c^4} \right)^{1/5}$ (4) $\left(\frac{K_{eq}}{25c^2} \right)^{1/5}$

Ans. NTA - (3)

Reso - (3)

Sol.

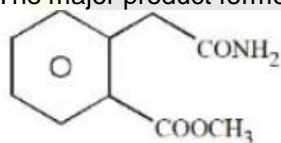


$$K = \frac{(2C\alpha)^2(3C\alpha)^3}{(C - C\alpha)}$$

$$K = \frac{108C^5\alpha^5}{C} = 108C^4\alpha^5$$

$$\alpha = \left[\frac{K}{108C^4} \right]^{1/5}$$

69. The major product formed in the following reaction is



- (1) (2) (3) (4)

Ans. NTA - (4)

Reso - (4)

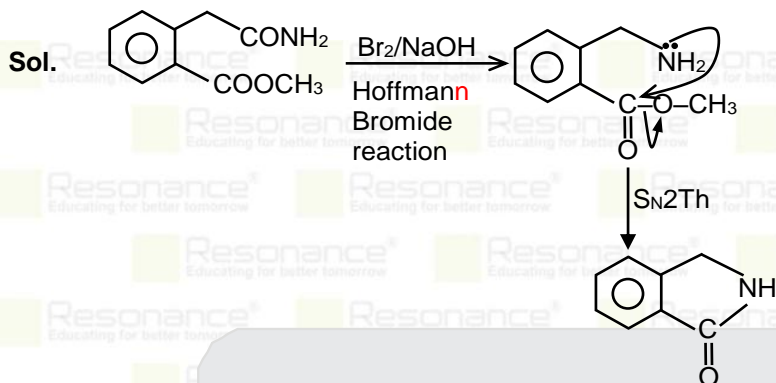
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70. The standard electrode potential of M^+/M in aqueous solution does not depend on
 (1) Ionisation of a gaseous metal atom (2) Ionisation of a solid metal atom
 (3) Hydration of a gaseous metal ion (4) Sublimation of a solid metal

Ans. NTA - (2)

Reso - (2)

Sol. The standard electrode potential of M^+/M in aqueous solution depend on

* Ionisation of a gaseous metal atom

* Hydration of a gaseous metal ion

* Sublimation of a solid metal

$E_{M^+/M}^0$ depends on

$M(s) \rightarrow M(g) ; \Delta H_{\text{sub}}$

$M(g) \rightarrow M^+(g) ; \Delta H_{\text{i.e.}}$

$M^+(g) + aq \rightarrow M^+(aq) ; \Delta H_{\text{Hydration}}$

71. Given below are two statements, one is labelled as Assertion A and the other is labelled as Reason R.

Assertion A : Loss of electron from hydrogen atom results in nucleus of $\sim 1.5 \times 10^{-3}$ pm size.

Reason R : Proton (H^+) always exists in combined form.

In the light of the above statements, choose the most appropriate answer from the options given below :

- (1) A is correct but R is not correct
 (2) A is not correct but R is correct
 (3) Both A and R are correct and R is the correct explanation of A
 (4) Both A and R are correct but R is NOT the correct explanation of A

Ans. NTA - (4)

Reso - (4)

Sol. Theory based

72. Match List I with List II :

	List I (Oxide)		List II (Type of bond)
(A)	N_2O_4	(I)	1 N = O bond
(B)	NO_2	(II)	1 N-O-N bond
(C)	N_2O_5	(III)	1 N-N bond
(D)	N_2O	(IV)	1 N=N/N≡N bond

Choose the correct answer from the options given below :

- (1) A - III, B - I, C - II, D - IV (2) A - III, B - I, C - IV, D - II
 (3) A - II, B - I, C - III, D - IV (4) A - II, B - IV, C - III, D - I

Ans. NTA - (1)






Reso - (1)

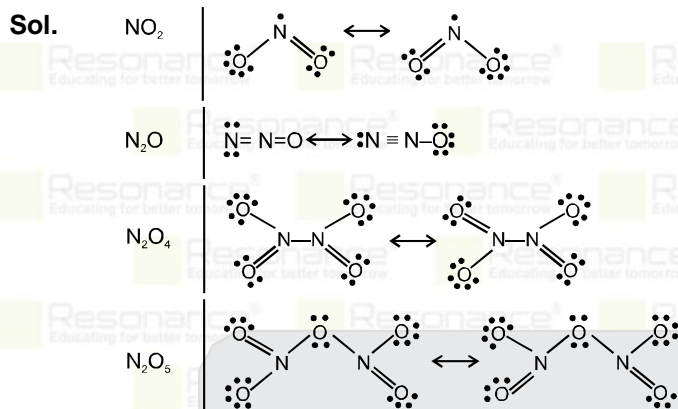
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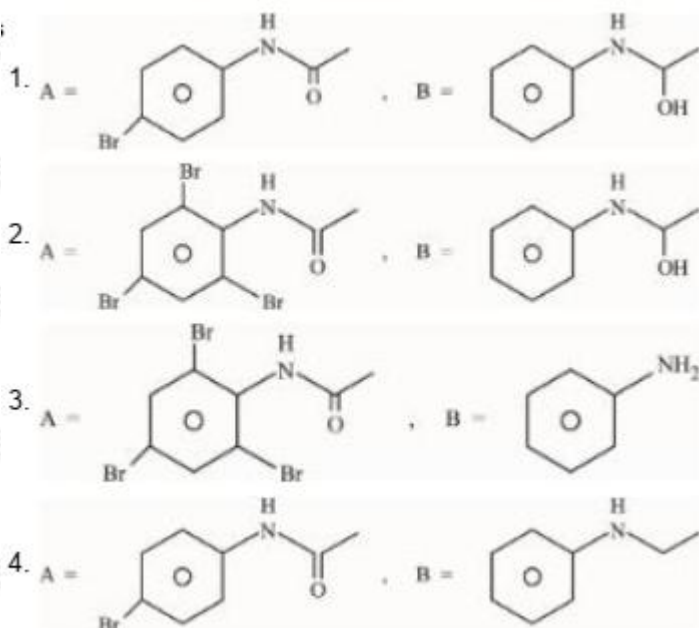
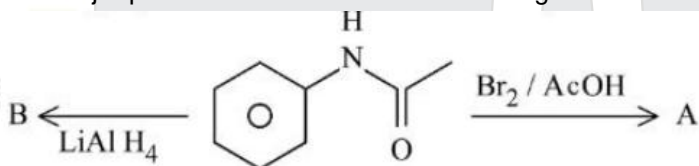
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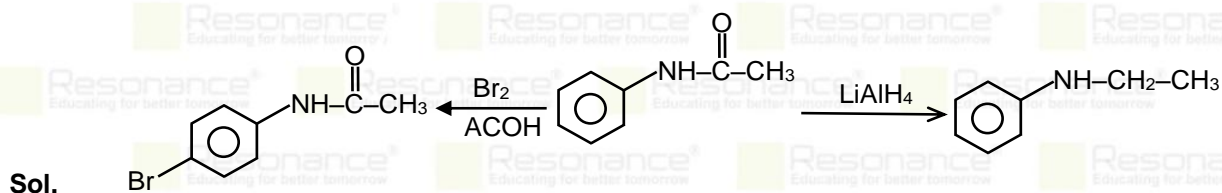
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73. The major products A and B from the following reactions are :



Ans. NTA - (4)
Reso - (4)



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74. Strong reducing and oxidizing agents among the following, respectively, are :

- (1) Ce^{4+} and Tb^{4+} (2) Ce^{4+} and Eu^{2+}
(3) Eu^{2+} and Ce^{4+} (4) Ce^{3+} and Ce^{4+}

Ans. NTA - (3)

Reso - (3)

Sol. $(E_{Ce^{+4}/Ce^{+3}}^0)_{RP} = 1.74 V$

Eu^{2+} is a strong reducing agent changing to common oxidation state + 3.

Most common oxidation no. of Eu & Ce are + 3

75. Match List I with List II :

	List I (Enzymatic reaction)		List II (Enzyme)
(A)	Sucrose \rightarrow Glucose and Fructose	(I)	Zymase
(B)	Glucose \rightarrow ethyl alcohol and CO_2	(II)	Pepsin
(C)	Starch \rightarrow Maltose	(III)	Invertase
(D)	Proteins \rightarrow Amino acids	(IV)	Diastase

Choose the correct answer from the options given below :

- (1) A - III, B - I, C - IV, D - II (2) A - III, B - I, C - II, D - IV
(3) A - I, B - IV, C - III, D - II (4) A - I, B - II, C - IV, D - III

Ans. NTA - (1)

Reso - (1)

Sol. Cane sugar (sucrose) $\xrightarrow{\text{Invertase}}$ Glucose + Fructose

Glucose $\xrightarrow{\text{Zymase}}$ ethyl alcohol

Starch $\xrightarrow{\text{Diastase}}$ Maltose

Protein $\xrightarrow{\text{Pepsin}}$ Amino acid

76. The possibility of photochemical smog formation is more at

- (1) The places with healthy vegetation (2) Marshy lands
(3) Industrial areas (4) Himalayan villages in winter

Ans. NTA - (3)

Reso - (3)

Sol. Fact based

77. The difference between electron gain enthalpies will be maximum between :

- (1) Ne and F (2) Ar and F (3) Ne and Cl (4) Ar and Cl

Ans. NTA - (3)

Reso - (3)

Sol. Cl has maximum -ve ΔH_{eg} and Ne has most +ive ΔH_{eg} hence difference will be maximum for Ne and Cl.

Element ΔH_{eg} (KJ/mole)

F -333

Cl -349

Ne +48

Ar +116

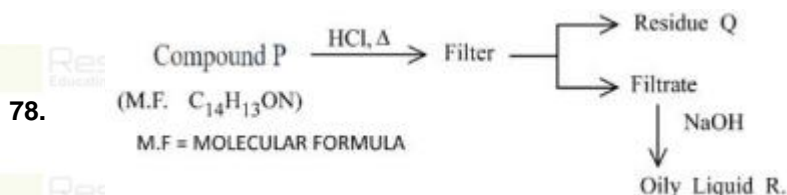
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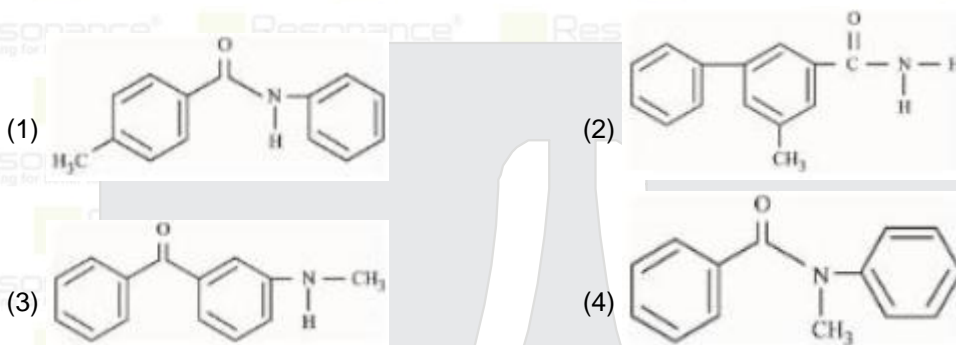
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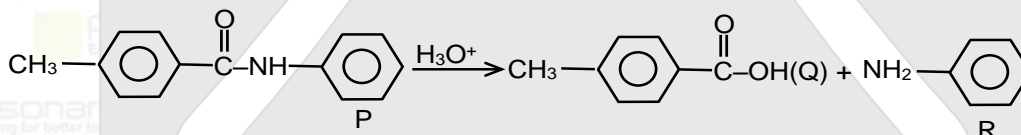
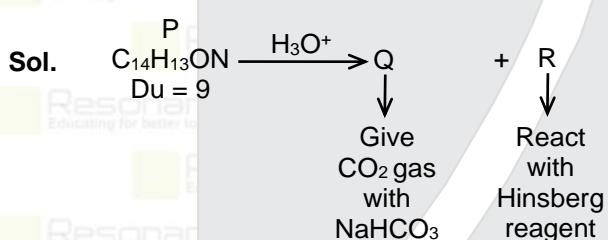


Compound P is neutral, Q gives effervescence with NaHCO_3 while R reacts with Hinsbergs reagent to give solid soluble in NaOH, compound P is

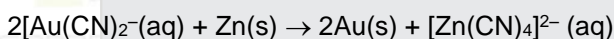


Ans. NTA - (1)

Reso - (1)



79. Which of the following options are correct for the reaction



(A) Redox reaction

(B) Displacement reaction

(C) Decomposition reaction

(D) Combination reaction

Choose the correct answer from the options given below :

(1) (A) and (B) only

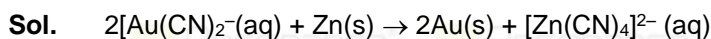
(2) (A) only

(3) (C) and (D) only

(4) (A) and (D) only

Ans. NTA - (1)

Reso - (1)



Above reaction is type of Redox reaction as well as Displacement reaction.

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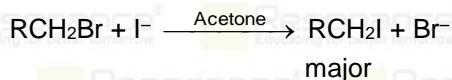
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80. For the reaction



The correct statement is

- (1) The solvent used in the reaction solvates the ions formed in rate determining step.
- (2) Br^- can act as competing nucleophile.
- (3) The transition state formed in the above reaction is less polar than the localised anion.
- (4) The reaction can occur in acetic acid also.

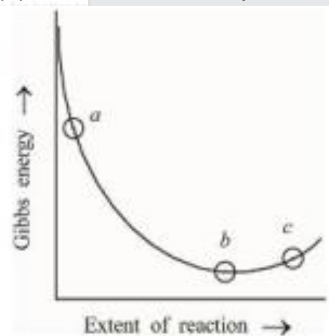
Ans. NTA - (3)

Reso - (3)

Sol. This is a $\text{S}_{\text{N}}2$ reaction complete by transition state which requires strong nucleophile in nonpolar/polar aprotic solvent.

81. Consider the graph of Gibbs free energy G vs Extent of reaction. The number of statement/s from the following which are true with respect to points (a), (b) and (c) is _____

- (1) Reaction is spontaneous at (a) and (b)
- (2) Reaction is at equilibrium at point (b) and non-spontaneous at point (c)
- (3) Reaction is spontaneous at (a) and non-spontaneous at (c)
- (4) Reaction is non-spontaneous at (a) and (b)



Ans. NTA - (2)

Reso - (2)

Sol. Reaction is spontaneous in a direction in which Gibbs free energy decreases. Hence 2 statements i.e. B & C are correct.

82. In ammonium – phosphomolybdate, the oxidation state of Mo is + _____

Ans. NTA - (6)

Reso - (6)

Sol. The Oxidation state of Mo in ammonium phosphomolybdate = + 6
(NH_4)₃PO₄.12MoO₃

83. Number of ambidentate ligands in a representative metal complex $[\text{M}(\text{en})(\text{SCN})_4]$ is _____.
[en = ethylenediamine]

Ans. NTA - (4)

Reso - (4)

Sol. SCN^- is an ambidentate ligand.

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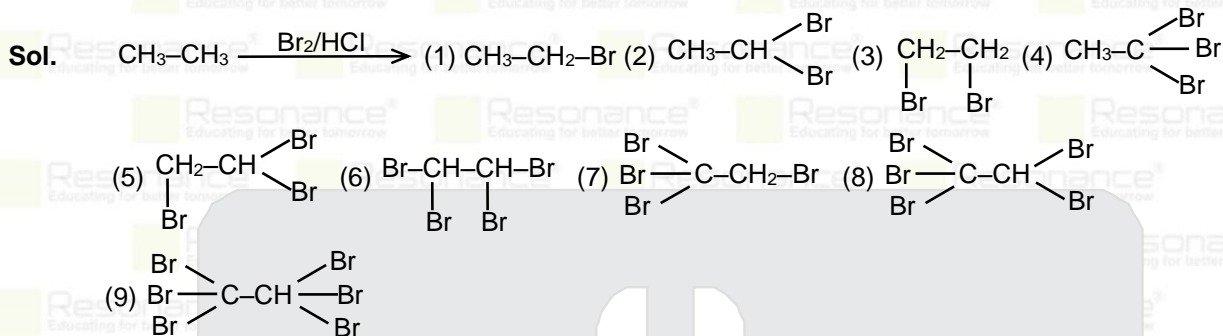
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84. Number of bromo derivatives obtained on treating ethane with excess of Br₂ in diffused sunlight is _____

Ans. NTA - (9)
Reso - (9)



85. Mass of Urea (NH₂CONH₂) required to be dissolved in 1000 g of water in order to reduce the vapour pressure of water by 25% is _____ g. (Nearest integer)

Given : Molar mass of N, C, O and H are 14, 12, 16 and 1 g mol⁻¹ respectively.

Ans. NTA - (1111)
Reso - (1111)

Sol. For a solution

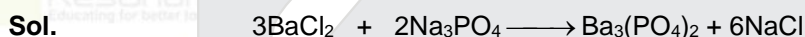
$$\frac{P^0 - P_s}{P_s} = \frac{n}{N}$$

$$\frac{100 - 75}{75} = \frac{w_{\text{urea}} \times 18}{60 \times 1000}$$

$$w_{\text{urea}} = 1111.1 \text{ gm}$$

86. If 5 moles of BaCl₂ is mixed with 2 moles of Na₃PO₄, the maximum number of moles of Ba₃(PO₄)₂ formed is _____ (Nearest integer)

Ans. NTA - (1)
Reso - (1)



Initial mole 5 2

$$\text{L.R.} \quad \frac{5}{3} = 1.67 \quad \frac{2}{2} = 1 \quad (\text{L.R.} = \text{Na}_3\text{PO}_4)$$

$$\frac{\text{mole Na}_3\text{PO}_4}{2} = \frac{\text{mole Ba}_3(\text{PO}_4)_2}{1}$$

$$\text{maximum number of mole of Ba}_3(\text{PO}_4)_2 = \frac{2}{2} = 1 \text{ mole}$$

87. The wavelength of an electron of kinetic energy 4.50×10^{-29} J is _____ $\times 10^{-5}$ m. (Nearest integer)

Given : mass of electron is 9×10^{-31} kg, $h = 6.6 \times 10^{-34}$ J s

Ans. NTA - (7)
Reso - (7)

Sol. $\lambda = \frac{h}{\sqrt{2m \text{ K.E.}}} = \frac{6.6 \times 10^{-34}}{\sqrt{2 \times 9.1 \times 10^{-31} \times 4.5 \times 10^{-29}}}$

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88. For the adsorption of hydrogen on platinum, the activation energy is 30 kJ mol^{-1} and for the adsorption of hydrogen on nickel, the activation energy is 41.4 kJ mol^{-1} . The logarithm of the ratio of the rates of chemisorption on equal areas of the metals at 300 K is _____ (Nearest integer)

$$R = 8.3 \text{ J K}^{-1} \text{ mol}^{-1}$$

Ans. NTA - (2)

Reso - (2)

Sol. $\ln \frac{K_1}{K_2} = \frac{E_{a_1} - E_{a_2}}{R.T.}$

$$\ln \frac{K_1}{K_2} = 2.$$

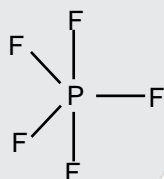
89. The number of species from the following which have square pyramidal structure is _____

$\text{PF}_5, \text{BrF}_4^-, \text{IF}_5, \text{BrF}_5, \text{XeOF}_4, \text{ICl}_4^-$

Ans. NTA - (3)

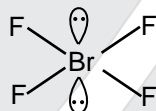
Reso - (3)

Sol. PF_5



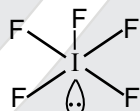
Triangular bipyramidal

BrF_4^-



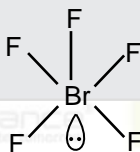
Square planar

IF_5



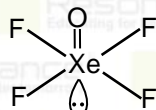
Square pyramidal

BrF_5



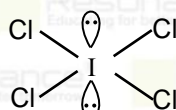
Square pyramidal

XeOF_4



Square pyramidal

ICl_4^-



Square planar

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90. The value of $\log K$ for the reaction $A \rightleftharpoons B$ at 298 K is _____. (Nearest integer)

Given : $\Delta H^\circ = -54.07 \text{ kJ mol}^{-1}$

$\Delta S^\circ = 10 \text{ J K}^{-1} \text{ mol}^{-1}$

(Take $2.303 \times 8.314 \times 298 = 5705$)

Ans. NTA - (10)

Reso - (10)

Sol. $\Delta G^\circ = \Delta H^\circ - T\Delta S^\circ$

$$\Delta G^\circ = -54070 - 293 \times 10 = -57000 \text{ J}$$

$$\Delta G^\circ = -2.303 RT \log K_{\text{eq}}$$

$$-57000 \text{ J} = -5705 \log K_{\text{eq}}$$






$$\log K_{\text{eq}} = 9.99 = 10$$

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