

**CODE****8****PAPER-1****P1-14-8****1514308****Time : 3 Hours****Maximum Marks : 180****Please read the instructions carefully. You are allotted 5 minutes specifically for this purpose.****INSTRUCTIONS****A. General**

1. This booklet is your Question Paper. Do not break the seal of this booklet before being instructed to do so by the invigilators.
2. The question paper CODE is printed on the left hand top corner of this sheet and on the back cover page of this booklet.
3. Blank spaces and blank pages are provided in the question paper for your rough work. No additional sheets will be provided for rough work.
4. Blank papers, clipboards, log tables, slide rules, calculators, cameras, cellular phones, pagers and electronic gadget of any kind are NOT allowed inside the examination hall.
5. Write your Name and Roll number in the space provided on the back cover of this booklet.
6. Answers to the questions and personal details are to be filled on an Optical Response Sheet, which is provided separately. The ORS is a doublet of two sheets - upper and lower, having identical layout. The upper sheet is a machine-gradable Objective Response Sheet (ORS) which will be collected by the invigilator at the end of the examination. The upper sheet is designed in such a way that darkening the bubble with a ball point pen will leave an identical impression at the corresponding place on the lower sheet. You will be allowed to take away the lower sheet at the end of the examination. (see Figure-1 on the back cover page for the correct way of darkening the bubbles for valid answers).
7. **Use a black ball point pen only to darken the bubbles on the upper original sheet.** Apply sufficient pressure so that the impression is created on the lower sheet. See Figure-1 on the back cover page for appropriate way of darkening the bubbles for valid answers.
8. **DO NOT TAMPER WITH / MUTILATE THE ORS OR THIS BOOKLET.**
9. On breaking the seal of the booklet check that it contains **28** pages and all the **60** questions and corresponding answer choices are legible. Read carefully the instruction printed at the beginning of each section.

**B. Filling the right part of the ORS**

10. The ORS also has a **CODE** printed on its left and right parts.
11. Verify that the CODE printed on the ORS (on both the left and right parts) is the same as that on this booklet and put your signature in the Box designated as R4.
12. **IF THE CODES DO NOT MATCH, ASK FOR A CHANGE OF THE BOOKLET / ORS AS APPLICABLE.**
13. Write your Name, Roll No. and the name of centre and sign with pen in the boxes provided on the upper sheet of ORS. **Do not write any of this anywhere else.** Darken the appropriate bubble **UNDER** each digit of your Roll No. in such way that the impression is created on the bottom sheet. (see example in Figure 2 on the back cover)

**C. Question Paper Format**

- The question paper consists of **three parts** (Physics, Chemistry and Mathematics). Each part consists of two sections.
14. **Section 1** contains **10 multiple choice questions**. Each question has four choices (A), (B), (C) and (D) out of which **ONE OR MORE THAN ONE** are correct.
  15. **Section 2** contains **10 questions**. The answer to each of the questions is a single-digit integer, ranging from 0 to 9 (both inclusive) .

**DO NOT BREAK THE SEAL WITHOUT BEING INSTRUCTED TO DO SO BY THE INVIGILATOR****Please read the last page of this booklet for rest of the instructions.**

	Subject	Section		Page No.
Part I	Physics	1	One or More Than One Option Correct Type	3 - 7
		2	One Integer Value Correct Type	8 - 12
Part II	Chemistry	1	One or More Than One Option Correct Type	13 - 17
		2	One Integer Value Correct Type	18 - 19
Part III	Mathematics	1	One or More Than One Option Correct Type	20 - 23
		2	One Integer Value Correct Type	24 - 26

Space for Rough Work

$$R = \frac{\rho \times L}{\pi \times d^2 / 4}$$

$$\frac{\rho \times L}{\pi \times d^2}$$

$$\frac{\rho \times L \times 4}{\pi \times d^2}$$

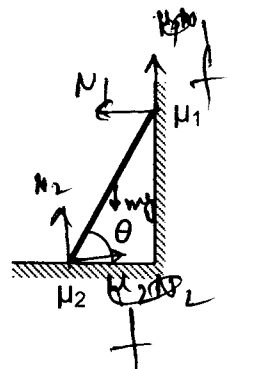
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3. In the figure, a ladder of mass  $m$  is shown leaning against a wall. It is in static equilibrium making an angle  $\theta$  with the horizontal floor. The coefficient of friction between the wall and the ladder is  $\mu_1$  and that between the floor and the ladder is  $\mu_2$ . The normal reaction of the wall on the ladder is  $N_1$  and that of the floor is  $N_2$ . If the ladder is about to slip, then

- (A)  $\mu_1 = 0$   $\mu_2 \neq 0$  and  $N_2 \tan \theta = \frac{mg}{2}$   
 (B)  $\mu_1 \neq 0$   $\mu_2 = 0$  and  $N_1 \tan \theta = \frac{mg}{2}$   
 (C)  $\mu_1 \neq 0$   $\mu_2 \neq 0$  and  $N_2 = \frac{mg}{1 + \mu_1 \mu_2}$   
 (D)  $\mu_1 = 0$   $\mu_2 \neq 0$  and  $N_1 \tan \theta = \frac{mg}{2}$



4. A light source, which emits two wavelengths  $\lambda_1 = 400 \text{ nm}$  and  $\lambda_2 = 600 \text{ nm}$ , is used in a Young's double slit experiment. If recorded fringe widths for  $\lambda_1$  and  $\lambda_2$  are  $\beta_1$  and  $\beta_2$  and the number of fringes for them within a distance  $y$  on one side of the central maximum are  $m_1$  and  $m_2$ , respectively, then
- (A)  $\beta_2 > \beta_1$   
 (B)  $m_1 > m_2$   
 (C) From the central maximum, 3<sup>rd</sup> maximum of  $\lambda_2$  overlaps with 5<sup>th</sup> minimum of  $\lambda_1$   
 (D) The angular separation of fringes for  $\lambda_1$  is greater than  $\lambda_2$

Space for Rough Work

CP

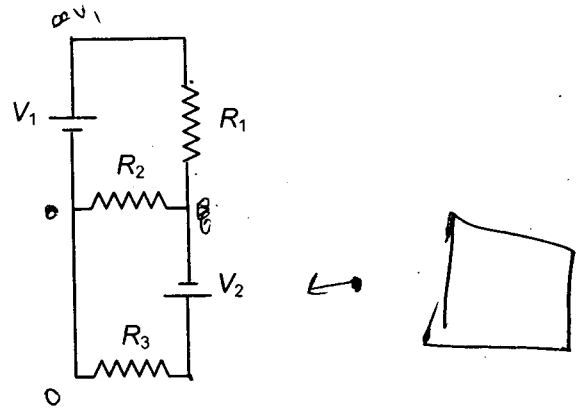
$$\beta_2 = \frac{\lambda_2 D}{d}$$

$$\beta_1 = \frac{\lambda_1 D}{d}$$



5. Two ideal batteries of emf  $V_1$  and  $V_2$  and three resistances  $R_1$ ,  $R_2$  and  $R_3$  are connected as shown in the figure. The current in resistance  $R_2$  would be zero if

- (A)  $V_1 = V_2$  and  $R_1 = R_2 = R_3$
- (B)  $V_1 = V_2$  and  $R_1 = 2R_2 = R_3$
- (C)  $V_1 = 2V_2$  and  $2R_1 = 2R_2 = R_3$
- (D)  $2V_1 = V_2$  and  $2R_1 = R_2 = R_3$



6. Let  $E_1(r)$ ,  $E_2(r)$  and  $E_3(r)$  be the respective electric fields at a distance  $r$  from a point charge  $Q$ , an infinitely long wire with constant linear charge density  $\lambda$ , and an infinite plane with uniform surface charge density  $\sigma$ . If  $E_1(r_0) = E_2(r_0) = E_3(r_0)$  at a given distance  $r_0$ , then

- (A)  $Q = 4\sigma\pi r_0^2$
- (B)  $r_0 = \frac{\lambda}{2\pi\sigma}$
- (C)  $E_1(r_0/2) = 2E_2(r_0/2)$
- (D)  $E_2(r_0/2) = 4E_3(r_0/2)$

Space for Rough Work

$kq/r^2$

$\frac{kQ}{r_0^2} = \frac{2k\lambda}{r_0} = \frac{G}{2\epsilon_0}$

$N_1 = N_2 N_2$

$N_2 + N_1 N_1 = mg$

$V_1 = i_1 R_1$

$\frac{4kQ}{r_0^2}$

(N<sub>1</sub>)

$\frac{kQ/r_0^2}{\frac{4kQ}{r_0^2}} = \frac{2 \times 2kQ/r_0^2}{\frac{4kQ}{r_0^2}}$

$\frac{r}{r_0^2} = \frac{2kQ}{r_0^2}$

$r_0 =$



$-V_1 + i_1 R_1 + i_3 R_3 = 0$

$-V_2 + i_2 R_2 - i_3 R_3 = 0$

$-V_1 + i_1 R_1 - V_2 + i_2 R_2 = 0$

$V_2 = i_2 R_2$

$V_2 + V_1 = i_1 R_1$

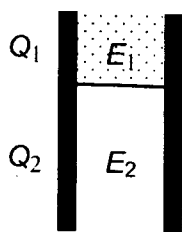


7. A student is performing an experiment using a resonance column and a tuning fork of frequency  $244 \text{ s}^{-1}$ . He is told that the air in the tube has been replaced by another gas (assume that the column remains filled with the gas). If the minimum height at which resonance occurs is  $(0.350 \pm 0.005) \text{ m}$ , the gas in the tube is

(Useful information :  $\sqrt{167RT} = 640 \text{ J}^{1/2} \text{ mole}^{-1/2}$ ;  $\sqrt{140RT} = 590 \text{ J}^{1/2} \text{ mole}^{-1/2}$ . The molar masses  $M$  in grams are given in the options. Take the values of  $\sqrt{\frac{10}{M}}$  for each gas as given there.)

- (A) Neon ( $M = 20, \sqrt{\frac{10}{20}} = \frac{7}{10}$ )      (B) Nitrogen ( $M = 28, \sqrt{\frac{10}{28}} = \frac{3}{5}$ )  
 (C) Oxygen ( $M = 32, \sqrt{\frac{10}{32}} = \frac{9}{16}$ )      (D) Argon ( $M = 36, \sqrt{\frac{10}{36}} = \frac{17}{32}$ )

8. A parallel plate capacitor has a dielectric slab of dielectric constant  $K$  between its plates that covers  $1/3$  of the area of its plates, as shown in the figure. The total capacitance of the capacitor is  $C$  while that of the portion with dielectric in between is  $C_1$ . When the capacitor is charged, the plate area covered by the dielectric gets charge  $Q_1$  and the rest of the area gets charge  $Q_2$ . The electric field in the dielectric is  $E_1$  and that in the other portion is  $E_2$ . Choose the correct option/options, ignoring edge effects.



- (A)  $\frac{E_1}{E_2} = 1$       (B)  $\frac{E_1}{E_2} = \frac{1}{K}$       (C)  $\frac{Q_1}{Q_2} = \frac{3}{K}$       (D)  $\frac{C}{C_1} = \frac{2+K}{K}$

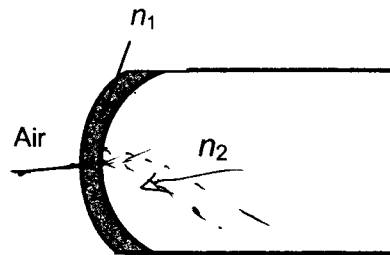
Space for Rough Work

$$C = \frac{C_0 A}{d - t(1 - \frac{1}{K})}$$

$$C_1 = \frac{C_0 A}{d - t(1 - \frac{1}{K})}$$

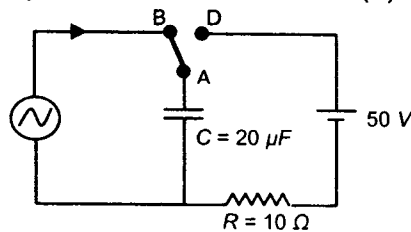


9. A transparent thin film of uniform thickness and refractive index  $n_1 = 1.4$  is coated on the convex spherical surface of radius  $R$  at one end of a long solid glass cylinder of refractive index  $n_2 = 1.5$ , as shown in the figure. Rays of light parallel to the axis of the cylinder traversing through the film from air to glass get focused at distance  $f_1$  from the film, while rays of light traversing from glass to air get focused at distance  $f_2$  from the film. Then



- (A)  ~~$|f_1| = 3R$~~
- (B)  $|f_1| = 2.8R$
- (C)  $|f_2| = 2R$
- (D)  ~~$|f_2| = 1.4R$~~

10. At time  $t = 0$ , terminal A in the circuit shown in the figure is connected to B by a key and an alternating current  $I(t) = I_0 \cos(\omega t)$ , with  $I_0 = 1A$  and  $\omega = 500 \text{ rad s}^{-1}$  starts flowing in it with the initial direction shown in the figure. At  $t = \frac{7\pi}{6\omega}$ , the key is switched from B to D. Now onwards only A and D are connected. A total charge  $Q$  flows from the battery to charge the capacitor fully. If  $C = 20 \mu F$ ,  $R = 10 \Omega$  and the battery is ideal with *emf* of 50V, identify the correct statement (s).



- (A) Magnitude of the maximum charge on the capacitor before  $t = \frac{7\pi}{6\omega}$  is  $1 \times 10^{-3} C$ .
- (B) The current in the left part of the circuit just before  $t = \frac{7\pi}{6\omega}$  is clockwise.
- (C) Immediately after A is connected to D, the current in  $R$  is 10A.
- (D)  $Q = 2 \times 10^{-3} C$ .

Space for Rough Work

2

6 ~~no~~ ~~no~~

no

~~no~~  $\frac{R}{2R} = \frac{2}{2}$

$\frac{4 \text{ KO}}{20}$

$\frac{8 \text{ KO}}{20}$

\* 8

7



## SECTION – 2 : (One Integer Value Correct Type)

This section contains 10 questions. Each question, when worked out will result in one integer from 0 to 9 (both inclusive).

11. Two parallel wires in the plane of the paper are distance  $X_0$  apart. A point charge is moving with speed  $u$  between the wires in the same plane at a distance  $X_1$  from one of the wires. When the wires carry current of magnitude  $I$  in the same direction, the radius of curvature of the path of the point charge is  $R_1$ . In contrast, if the currents  $I$  in the two wires have directions opposite to each other, the radius of curvature of the path is  $R_2$ . If  $\frac{X_0}{X_1} = 3$ , the value of  $\frac{R_1}{R_2}$  is
12. During Searle's experiment, zero of the Vernier scale lies between  $3.20 \times 10^{-2} m$  and  $3.25 \times 10^{-2} m$  of the main scale. The 20<sup>th</sup> division of the Vernier scale exactly coincides with one of the main scale divisions. When an additional load of 2 kg is applied to the wire, the zero of the Vernier scale still lies between  $3.20 \times 10^{-2} m$  and  $3.25 \times 10^{-2} m$  of the main scale but now the 45<sup>th</sup> division of Vernier scale coincides with one of the main scale divisions. The length of the thin metallic wire is 2 m and its cross-sectional area is  $8 \times 10^{-7} m^2$ . The least count of the Vernier scale is  $1.0 \times 10^{-5} m$ . The maximum percentage error in the Young's modulus of the wire is

Space for Rough Work

$$\frac{kq}{r^2} = \frac{2k\lambda}{r}$$

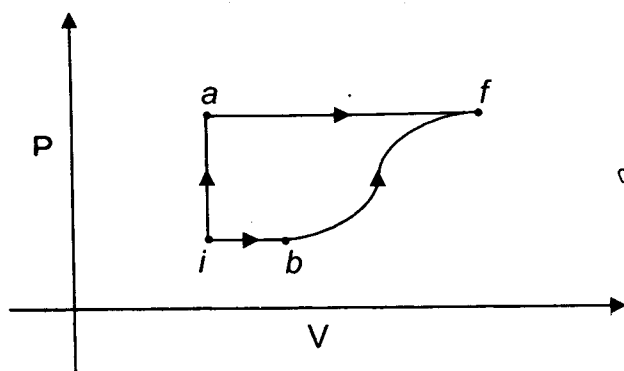
$$\frac{q}{4\pi\epsilon_0 r^2} = \frac{2\lambda}{4\pi\epsilon_0 r} = \frac{6}{2\epsilon_0}$$





13. To find the distance  $d$  over which a signal can be seen clearly in foggy conditions, a railways engineer uses dimensional analysis and assumes that the distance depends on the mass density  $\rho$  of the fog, intensity (power/area)  $S$  of the light from the signal and its frequency  $f$ . The engineer finds that  $d$  is proportional to  $S^{1/n}$ . The value of  $n$  is

14. A thermodynamic system is taken from an initial state  $i$  with internal energy  $U_i = 100\text{ J}$  to the final state  $f$  along two different paths  $iaf$  and  $ibf$ , as schematically shown in the figure. The work done by the system along the paths  $af$ ,  $ib$  and  $bf$  are  $W_{af} = 200\text{ J}$ ,  $W_{ib} = 50\text{ J}$  and  $W_{bf} = 100\text{ J}$  respectively. The heat supplied to the system along the path  $iaf$ ,  $ib$  and  $bf$  are  $Q_{iaf}$ ,  $Q_{ib}$  and  $Q_{bf}$  respectively. If the internal energy of the system in the state  $b$  is  $U_b = 200\text{ J}$  and  $Q_{iaf} = 500\text{ J}$ , the ratio  $Q_{bf}/Q_{ib}$  is



Handwritten calculations for question 14:

$$\frac{2K\lambda}{910} = \frac{6}{200}$$

$$\frac{2K\lambda}{4TTG \times 910} \neq$$

$$\frac{6 \times 910 \times 4TTG}{200}$$

Space for Rough Work

~~Handwritten scribble~~

Handwritten calculations for question 14:

$$Q = \frac{2 \times 2K\lambda \times 200}{910}$$

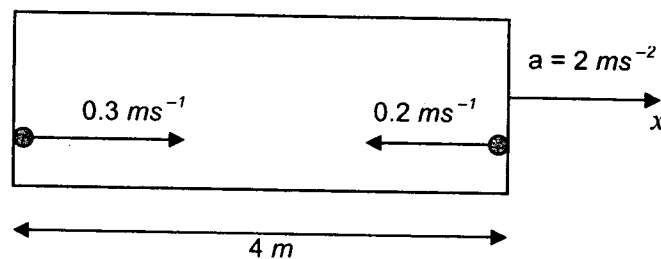
$$\frac{8K\lambda}{910}$$

$$\frac{4K\lambda}{910}$$

$$\frac{4 \times 6}{200}$$



15. A galvanometer gives full scale deflection with  $0.006 \text{ A}$  current. By connecting it to a  $4990 \Omega$  resistance, it can be converted into a voltmeter of range  $0 - 30 \text{ V}$ . If connected to a  $\frac{2n}{249} \Omega$  resistance, it becomes an ammeter of range  $0 - 1.5 \text{ A}$ . The value of  $n$  is
16. A rocket is moving in a gravity free space with a constant acceleration of  $2 \text{ ms}^{-2}$  along  $+x$  direction (see figure). The length of a chamber inside the rocket is  $4 \text{ m}$ . A ball is thrown from the left end of the chamber in  $+x$  direction with a speed of  $0.3 \text{ ms}^{-1}$  relative to the rocket. At the same time, another ball is thrown in  $-x$  direction with a speed of  $0.2 \text{ ms}^{-1}$  from its right end relative to the rocket. The time in seconds when the two balls hit each other is

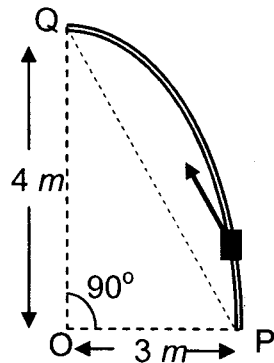


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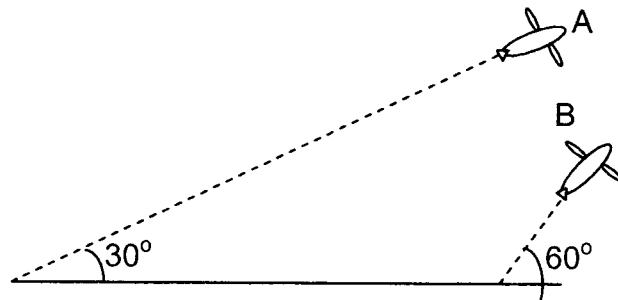




19. Consider an elliptically shaped rail PQ in the vertical plane with  $OP = 3\text{ m}$  and  $OQ = 4\text{ m}$ . A block of mass  $1\text{ kg}$  is pulled along the rail from P to Q with a force of  $18\text{ N}$ , which is always parallel to line PQ (see the figure given). Assuming no frictional losses, the kinetic energy of the block when it reaches Q is  $(n \times 10)\text{ Joules}$ . The value of  $n$  is (take acceleration due to gravity  $= 10\text{ ms}^{-2}$ )



20. Airplanes A and B are flying with constant velocity in the same vertical plane at angles  $30^\circ$  and  $60^\circ$  with respect to the horizontal respectively as shown in figure. The speed of A is  $100\sqrt{3}\text{ ms}^{-1}$ . At time  $t = 0\text{ s}$ , an observer in A finds B at a distance of  $500\text{ m}$ . This observer sees B moving with a constant velocity perpendicular to the line of motion of A. If at  $t = t_0$ , A just escapes being hit by B,  $t_0$  in seconds is



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*cosθ*

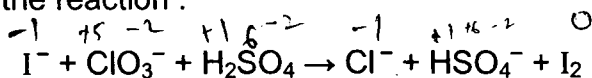


# PART II : CHEMISTRY

## SECTION – 1 : (One or More Than One Options Correct Type)

This section contains 10 multiple choice type questions. Each question has four choices (A), (B), (C) and (D) out of which ONE or MORE THAN ONE are correct.

21. For the reaction :



The correct statement(s) in the balanced equation is/are :

(A) Stoichiometric coefficient of  $\text{HSO}_4^-$  is 6.

~~(B) Iodide is oxidized.~~

(C) Sulphur is reduced.

~~(D)  $\text{H}_2\text{O}$  is one of the products.~~

$$1 + x - 8 = -1$$

$$2 - 8 + x = 0$$

$$x = 6$$

22. The pair(s) of reagents that yield paramagnetic species is/are

(A) Na and excess of  $\text{NH}_3$

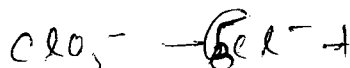
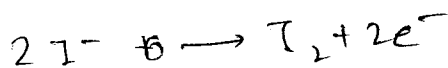
~~(B) K and excess of  $\text{O}_2$~~

~~(C) Cu and dilute  $\text{HNO}_3$~~

(D)  $\text{O}_2$  and 2-ethylantraquinol

$$x - 6 = -1$$

Space for Rough Work



$$\frac{2x}{6} = \frac{6}{2x}$$

$$\frac{2x}{6} = \frac{2x}{6} \times 2$$

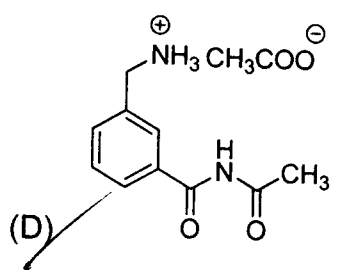
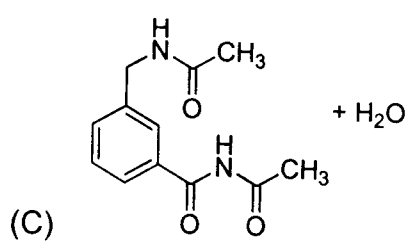
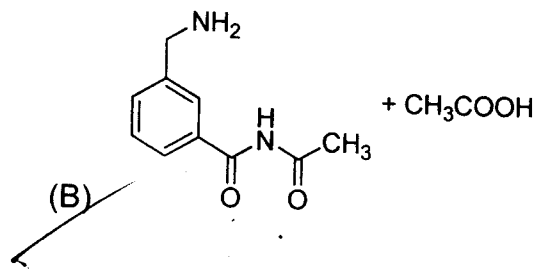
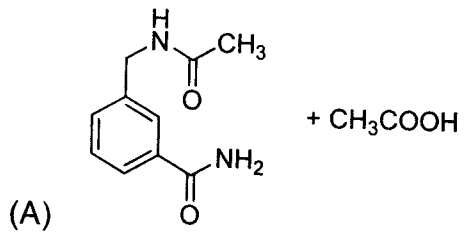
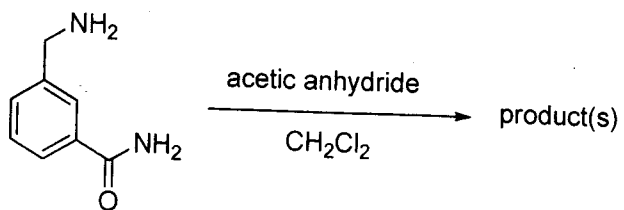
$$= \frac{2x}{6} \times 2$$

$$= \frac{4x}{6}$$



CHEMISTRY

23. In the reaction shown below, the major product(s) formed is/are



Space for Rough Work



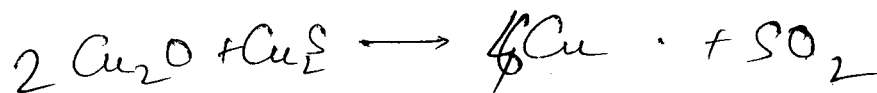
24. In a galvanic cell, the salt bridge

- ~~(A)~~ does not participate chemically in the cell reaction.  
~~(B)~~ stops the diffusion of ions from one electrode to another.  
~~(C)~~ is necessary for the occurrence of the cell reaction.  
(D) ensures mixing of the two electrolytic solutions.

25. Upon heating with  $\text{Cu}_2\text{S}$ , the reagent(s) that give copper metal is/are

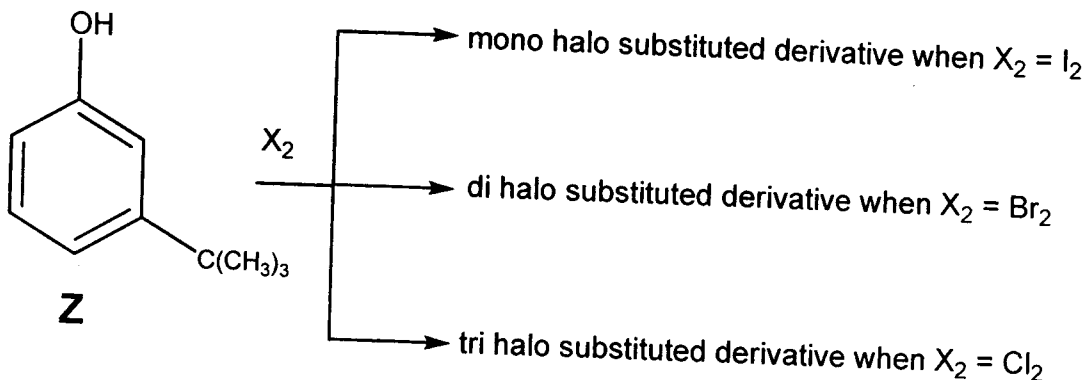
- ~~(A)~~  $\text{CuFeS}_2$                       ~~(B)~~  $\text{CuO}$   
(C)  $\text{Cu}_2\text{O}$                         (D)  $\text{CuSO}_4$

Space for Rough Work



26. Hydrogen bonding plays a central role in the following phenomena:
- (A) Ice floats in water.
  - (B) Higher Lewis basicity of primary amines than tertiary amines in aqueous solutions.
  - (C) Formic acid is more acidic than acetic acid.
  - (D) Dimerisation of acetic acid in benzene.

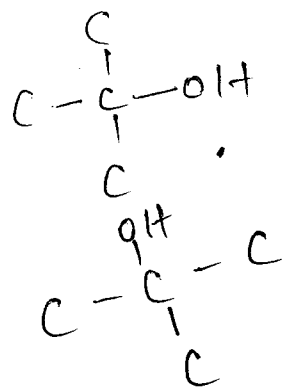
27. The reactivity of compound Z with different halogens under appropriate conditions is given below :



The observed pattern of electrophilic substitution can be explained by

- (A) the steric effect of the halogen
- (B) the steric effect of the *tert*-butyl group
- (C) the electronic effect of the phenolic group
- (D) the electronic effect of the *tert*-butyl group

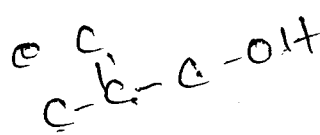
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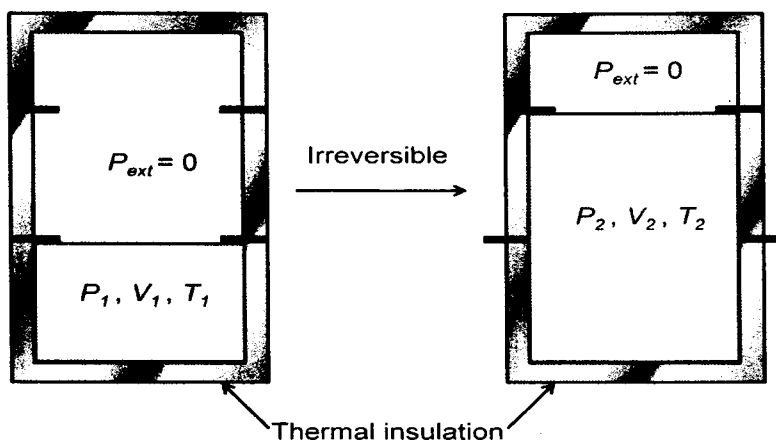


28. The correct combination of names for isomeric alcohols with molecular formula  $C_4H_{10}O$  is/are

- (A) *tert*-butanol and 2-methylpropan-2-ol ✓
- ~~(B) *tert*-butanol and 1, 1-dimethylethan-1-ol~~
- (C) *n*-butanol and butan-1-ol
- ~~(D) isobutyl alcohol and 2-methylpropan-1-ol~~



29. An ideal gas in a thermally insulated vessel at internal pressure =  $P_1$ , volume =  $V_1$  and absolute temperature =  $T_1$  expands irreversibly against zero external pressure, as shown in the diagram. The final internal pressure, volume and absolute temperature of the gas are  $P_2$ ,  $V_2$  and  $T_2$ , respectively. For this expansion,



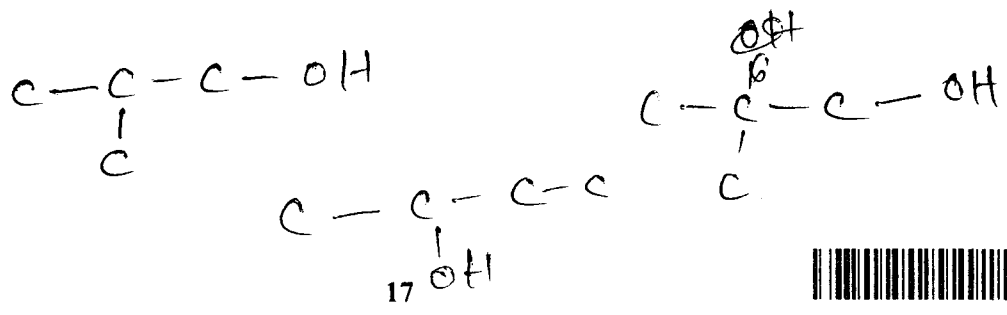
- (A)  $q = 0$
- ~~(B)  $T_2 = T_1$~~
- ~~(C)  $P_2V_2 = P_1V_1$~~
- (D)  $P_2V_2^\gamma = P_1V_1^\gamma$

25

30. The correct statement(s) for orthoboric acid is/are

- ~~(A) It behaves as a weak acid in water due to self ionization.~~
- (B) Acidity of its aqueous solution increases upon addition of ethylene glycol.
- ~~(C) It has a three dimensional structure due to hydrogen bonding.~~
- (D) It is a weak electrolyte in water.

Space for Rough Work

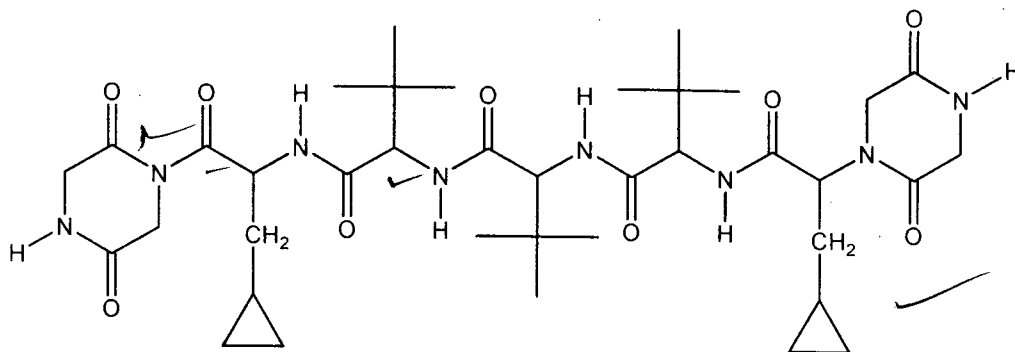


CHEMISTRY

SECTION – 2 : (One Integer Value Correct Type)

This section contains 10 questions. Each question, when worked out will result in one integer from 0 to 9 (both inclusive).

31. In an atom, the total number of electrons having quantum numbers  $n = 4$ ,  $|m_l| = 1$  and  $m_s = -1/2$  is 6
32. The total number of distinct naturally occurring amino acids obtained by complete acidic hydrolysis of the peptide shown below is 4



33. If the value of Avogadro number is  $6.023 \times 10^{23} \text{ mol}^{-1}$  and the value of Boltzmann constant is  $1.380 \times 10^{-23} \text{ J K}^{-1}$ , then the number of significant digits in the calculated value of the universal gas constant is 4
34. A compound  $\text{H}_2\text{X}$  with molar weight of 80 g is dissolved in a solvent having density of  $0.4 \text{ g ml}^{-1}$ . Assuming no change in volume upon dissolution, the **molality** of a 3.2 molar solution is 8
35.  $\text{MX}_2$  dissociates into  $\text{M}^{2+}$  and  $\text{X}^-$  ions in an aqueous solution, with a degree of dissociation ( $\alpha$ ) of 0.5. The ratio of the observed depression of freezing point of the aqueous solution to the value of the depression of freezing point in the absence of ionic dissociation is 3

Space for Rough Work

Handwritten calculations for question 34:

$$2 + x = 80$$

$$x = 78$$

$$\frac{x}{80} \times 0.4 = 3.2 \text{ t}$$

$$\frac{x}{80} \times 0.4 = 3.2$$

$$\frac{x}{80} = 3.2 \times \frac{10}{4} = 8$$

Handwritten calculations for question 35:

$$\frac{\Delta T_{\text{obs}}}{\Delta T_{\text{calc}}} = \frac{1 - \alpha + \alpha n}{1 - \alpha + \alpha n}$$

$$\frac{\Delta T_{\text{obs}}}{\Delta T_{\text{calc}}} = \frac{1 - 0.5 + 0.5 \times 3}{1 - 0.5 + 0.5 \times 2} = \frac{1.5}{2} = 0.75$$

Handwritten calculations for question 31:

$$n = 4, |m_l| = 1, m_s = -1/2$$

$$\text{Number of orbitals} = 2 \times n^2 \times |m_l| = 2 \times 4^2 \times 1 = 32$$

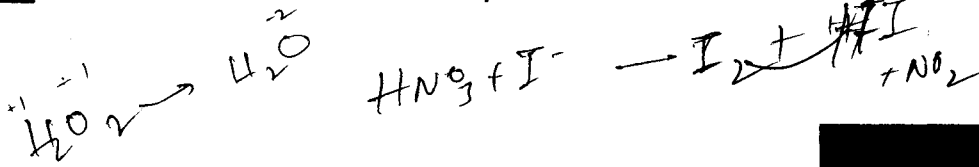
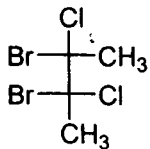
$$\text{Number of electrons} = 32 \times \frac{1}{2} = 16$$


36. Consider the following list of reagents :

Acidified  $K_2Cr_2O_7$ , alkaline  $KMnO_4$ ,  $CuSO_4$ ,  $H_2O_2$ ,  $Cl_2$ ,  $O_3$ ,  $FeCl_3$ ,  $HNO_3$  and  $Na_2S_2O_3$ .

The total number of reagents that can oxidise aqueous iodide to iodine is

37. The total number(s) of **stable** conformers with **non-zero** dipole moment for the following compound is (are)



38. Among  $PbS$ ,  $CuS$ ,  $HgS$ ,  $MnS$ ,  $Ag_2S$ ,  $NiS$ ,  $CoS$ ,  $Bi_2S_3$  and  $SnS_2$ , the total number of **BLACK** coloured sulphides is

39. Consider all possible isomeric ketones, including stereoisomers of MW = 100. All these isomers are independently reacted with  $NaBH_4$  (**NOTE**: stereoisomers are also reacted separately). The total number of ketones that give a racemic product(s) is/are

40. A list of species having the formula  $XZ_4$  is given below.

$XeF_4$ ,  $SF_4$ ,  $SiF_4$ ,  $BF_4^-$ ,  $BrF_4^-$ ,  $[Cu(NH_3)_4]^{2+}$ ,  $[FeCl_4]^{2-}$ ,  $[CoCl_4]^{2-}$  and  $[PtCl_4]^{2-}$ .

Defining shape on the basis of the location of X and Z atoms, the total number of species having a square planar shape is

CHEMISTRY

Space for Rough Work

Handwritten rough work for Q40:

- Species list:  $XeF_4$ ,  $SF_4$ ,  $SiF_4$ ,  $BF_4^-$ ,  $BrF_4^-$ ,  $[Cu(NH_3)_4]^{2+}$ ,  $[FeCl_4]^{2-}$ ,  $[CoCl_4]^{2-}$ ,  $[PtCl_4]^{2-}$
- Analysis of species:
  - $XeF_4$ : square planar
  - $SF_4$ : square pyramidal
  - $SiF_4$ : tetrahedral
  - $BF_4^-$ : tetrahedral
  - $BrF_4^-$ : square planar
  - $[Cu(NH_3)_4]^{2+}$ : square planar
  - $[FeCl_4]^{2-}$ : tetrahedral
  - $[CoCl_4]^{2-}$ : tetrahedral
  - $[PtCl_4]^{2-}$ : square planar
- Count of square planar species: 4 ( $XeF_4$ ,  $BrF_4^-$ ,  $[Cu(NH_3)_4]^{2+}$ ,  $[PtCl_4]^{2-}$ )
- Handwritten calculations:
  - $n \rightarrow 4$ ,  $l \rightarrow 0, 1, 2, 3$
  - $dsp^2$  hybridization
  - $0.5 = 2 - 1$
  - $1.5 = 1$
  - $\alpha = \frac{l-1}{n-1}$
  - $4s^1 3d^{10}$



# PART III : MATHEMATICS

## SECTION – 1 : (One or More Than One Options Correct Type)

This section contains 10 multiple choice questions. Each question has four choices (A), (B), (C) and (D) out of which ONE or MORE THAN ONE are correct.

41. Let  $f: [a, b] \rightarrow [1, \infty)$  be a continuous function and let  $g: \mathbb{R} \rightarrow \mathbb{R}$  be defined as

$$g(x) = \begin{cases} 0 & \text{if } x < a, \\ \int_a^x f(t) dt & \text{if } a \leq x \leq b, \\ \int_a^b f(t) dt & \text{if } x > b. \end{cases}$$

Then

- (A)  $g(x)$  is continuous but not differentiable at  $a$
- (B)  $g(x)$  is differentiable on  $\mathbb{R}$
- (C)  $g(x)$  is continuous but not differentiable at  $b$
- (D)  $g(x)$  is continuous and differentiable at either  $a$  or  $b$  but not both

42. For every pair of continuous functions  $f, g: [0, 1] \rightarrow \mathbb{R}$  such that

$$\max \{f(x): x \in [0, 1]\} = \max \{g(x): x \in [0, 1]\},$$

the correct statement(s) is(are) :

- (A)  $(f(c))^2 + 3f(c) = (g(c))^2 + 3g(c)$  for some  $c \in [0, 1]$
- (B)  $(f(c))^2 + f(c) = (g(c))^2 + 3g(c)$  for some  $c \in [0, 1]$
- (C)  $(f(c))^2 + 3f(c) = (g(c))^2 + g(c)$  for some  $c \in [0, 1]$
- (D)  $(f(c))^2 = (g(c))^2$  for some  $c \in [0, 1]$

Space for Rough Work

$[0, 1] \rightarrow \mathbb{R}$       $f$

$g(x) = \int_a^x f(t) dt$

$f\{t+3\} = g\{g+3\}$

$f\{t+1\} =$





46. Let  $M$  and  $N$  be two  $3 \times 3$  matrices such that  $MN = NM$ . Further, if  $M \neq N^2$  and  $M^2 = N^4$ , then
- (A) determinant of  $(M^2 + MN^2)$  is 0
  - (B) there is a  $3 \times 3$  non-zero matrix  $U$  such that  $(M^2 + MN^2)U$  is the zero matrix
  - (C) determinant of  $(M^2 + MN^2) \geq 1$
  - (D) for a  $3 \times 3$  matrix  $U$ , if  $(M^2 + MN^2)U$  equals the zero matrix then  $U$  is the zero matrix

47. Let  $f: (0, \infty) \rightarrow \mathbb{R}$  be given by

$$f(x) = \int_{\frac{1}{x}}^x e^{-(t+\frac{1}{t})} \frac{dt}{t}.$$

Then

- (A)  $f(x)$  is monotonically increasing on  $[1, \infty)$
- (B)  $f(x)$  is monotonically decreasing on  $(0, 1)$
- (C)  $f(x) + f\left(\frac{1}{x}\right) = 0$ , for all  $x \in (0, \infty)$
- (D)  $f(2^x)$  is an odd function of  $x$  on  $\mathbb{R}$

Space for Rough Work

$$\int_{\frac{1}{x}}^x \frac{e^{-(t+\frac{1}{t})}}{t} dt$$

LATE

$$= \frac{1}{t} e^t$$



48. Let  $f: (-\frac{\pi}{2}, \frac{\pi}{2}) \rightarrow \mathbb{R}$  be given by

$$f(x) = (\log(\sec x + \tan x))^3.$$

Then

- (A)  $f(x)$  is an odd function  
 (B)  $f(x)$  is a one-one function  
 (C)  $f(x)$  is an onto function  
 (D)  $f(x)$  is an even function

49. A circle  $S$  passes through the point  $(0, 1)$  and is orthogonal to the circles  $(x - 1)^2 + y^2 = 16$  and  $x^2 + y^2 = 1$ . Then

- (A) radius of  $S$  is 8  
 (B) radius of  $S$  is 7  
 (C) centre of  $S$  is  $(-7, 1)$   
 (D) centre of  $S$  is  $(-8, 1)$

50. Let  $a \in \mathbb{R}$  and let  $f: \mathbb{R} \rightarrow \mathbb{R}$  be given by

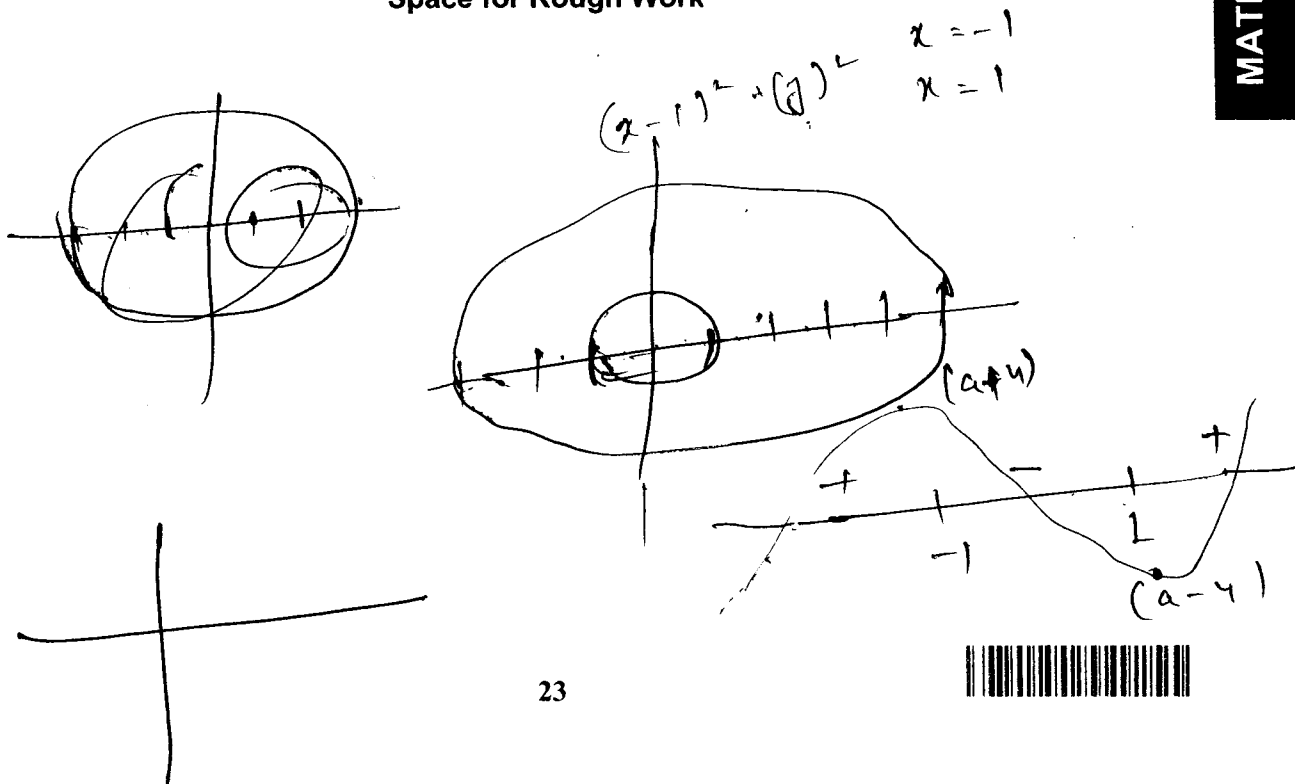
$$f(x) = x^5 - 5x + a.$$

Then

- (A)  $f(x)$  has three real roots if  $a > 4$   
 (B)  $f(x)$  has only one real root if  $a > 4$   
 (C)  $f(x)$  has three real roots if  $a < -4$   
 (D)  $f(x)$  has three real roots if  $-4 < a < 4$

Handwritten work for Q50:  
 $(1)^5 - 5 + a = 0 \Rightarrow a - 4 = 0 \Rightarrow a = 4$   
 $(-1)^5 - 5 + a = 0 \Rightarrow -1 - 5 + a = 0 \Rightarrow a = 6$   
 $5x^4 - 5 = 0 \Rightarrow x^4 = 1 \Rightarrow x = \pm 1, \pm i$   
 $5(x^2 + 1)(x^2 - 1)$   
 $5(x^2 + 1)(x + 1)(x - 1)$

Space for Rough Work



SECTION - 2 : (One Integer Value Correct Type)

This section contains 10 questions. Each question, when worked out will result in one integer from 0 to 9 (both inclusive).

51. The slope of the tangent to the curve  $(y - x^5)^2 = x(1 + x^2)^2$  at the point (1, 3) is

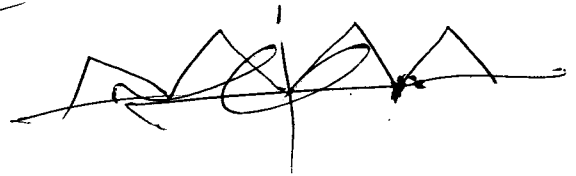
8

52. Let  $f: [0, 4\pi] \rightarrow [0, \pi]$  be defined by  $f(x) = \cos^{-1}(\cos x)$ . The number of points  $x \in [0, 4\pi]$  satisfying the equation

$f(x) = \frac{10-x}{10}$

is

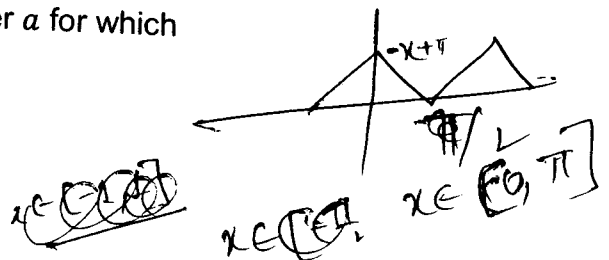
$x \in [-1, 1]$



53. The largest value of the non-negative integer  $a$  for which

$\lim_{x \rightarrow 1} \left\{ \frac{-ax + \sin(x-1) + a}{x + \sin(x-1) - 1} \right\}^{\frac{1-x}{1-\sqrt{x}}} = \frac{1}{4}$

is



54. Let  $f: \mathbb{R} \rightarrow \mathbb{R}$  and  $g: \mathbb{R} \rightarrow \mathbb{R}$  be respectively given by  $f(x) = |x| + 1$  and  $g(x) = x^2 + 1$ . Define  $h: \mathbb{R} \rightarrow \mathbb{R}$  by

$h(x) = \begin{cases} \max \{f(x), g(x)\} & \text{if } x \leq 0, \\ \min \{f(x), g(x)\} & \text{if } x > 0. \end{cases}$

$1 - \frac{x}{10}$

The number of points at which  $h(x)$  is not differentiable is

Space for Rough Work

$2(y - x^5)(y' - 5x^4) = x \{ 2(1+x^2)(2x) \}' + (1+x^2)^2$

$2(3-1)(y'-5) = 1 \{ 2(1+1)(2) \} + (1+1)^2$

$2(2)(y'-5) = \{ 2(2)(2) \} + (2)^2$

$4(y'-5) = 4 + 4$

$y' = 8$

\* 8

0 =





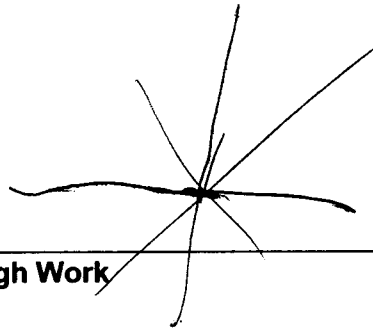
55. For a point  $P$  in the plane, let  $d_1(P)$  and  $d_2(P)$  be the distances of the point  $P$  from the lines  $x - y = 0$  and  $x + y = 0$  respectively. The area of the region  $R$  consisting of all points  $P$  lying in the first quadrant of the plane and satisfying  $2 \leq d_1(P) + d_2(P) \leq 4$ , is

56. Let  $n_1 < n_2 < n_3 < n_4 < n_5$  be positive integers such that  $n_1 + n_2 + n_3 + n_4 + n_5 = 20$ . Then the number of such distinct arrangements  $(n_1, n_2, n_3, n_4, n_5)$  is

57. The value of

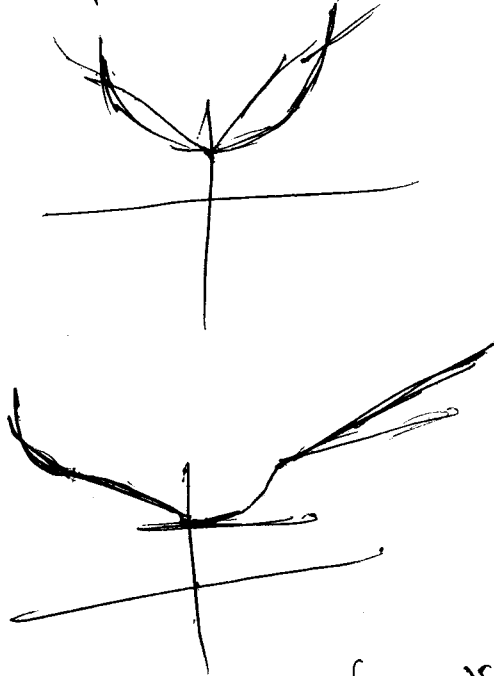
$$\int_0^1 4x^3 \left\{ \frac{d^2}{dx^2} (1-x^2)^5 \right\} dx$$

is



Space for Rough Work

$$2 \leq d_1(P) + d_2(P) \leq 4$$



$$4x^3 \left\{ (1-x^2)^5 \right\}' - \int (1-x^2)^5 \cdot 12x^2$$

$$4x^3 (y') - \left\{ 12x^2 (1-x^2)^5 - \int (1-x^2)^5 (22x) \right\}$$



58. Let  $\vec{a}$ ,  $\vec{b}$ , and  $\vec{c}$  be three non-coplanar unit vectors such that the angle between every pair of them is  $\frac{\pi}{3}$ . If  $\vec{a} \times \vec{b} + \vec{b} \times \vec{c} = p\vec{a} + q\vec{b} + r\vec{c}$ , where  $p, q$  and  $r$  are scalars, then the value of  $\frac{p^2 + 2q^2 + r^2}{q^2}$  is

59. Let  $a, b, c$  be positive integers such that  $\frac{b}{a}$  is an integer. If  $a, b, c$  are in geometric progression and the arithmetic mean of  $a, b, c$  is  $b + 2$ , then the value of

$$\frac{a^2 + a - 14}{a + 1}$$

is

60. Let  $n \geq 2$  be an integer. Take  $n$  distinct points on a circle and join each pair of points by a line segment. Colour the line segment joining every pair of adjacent points by blue and the rest by red. If the number of red and blue line segments are equal, then the value of  $n$  is

Space for Rough Work

$b = ak$

$b^2 = ac$

$a^2 k^2 = ac$

$k^2 = \frac{c}{a}$

$\frac{a+c}{2} = b+2$

$a+c = 2b+4$

$a+ak^2 = 2(ak)+4$

$a(k^2+1-2k) = 4$

$a(k-1)^2 = 4$

$= 1 \times 4$

$=$

$\frac{1+1-14}{2}$

$b^2 = ac$

$a^2 k^2 = ac$

$k^2 = \frac{c}{a}$

$\frac{16+4-14}{5}$

$\frac{1+1-14}{2}$

$\frac{16+4-14}{5}$

$\frac{a^2+a-14}{a+1}$

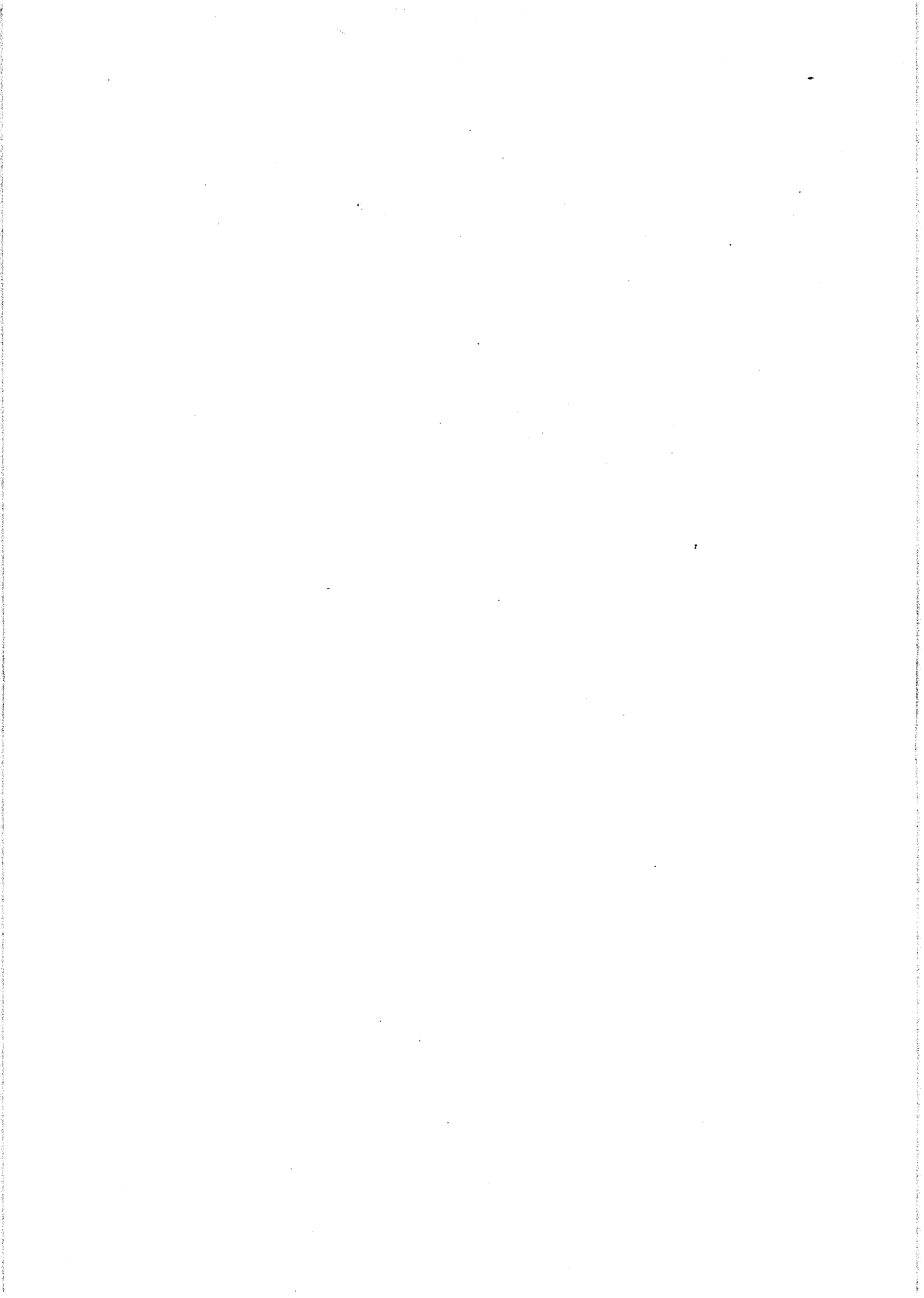
$20-14$

$\frac{1+1-14}{2}$

$\frac{16+4-14}{5}$

$\frac{a^2+a-14}{a+1}$





**D. Marking Scheme**

16. For each question in **Section 1**, you will be awarded **3 marks** if you darken all the bubble(s) corresponding to the correct answer(s) and **zero mark** if no bubbles are darkened. **No negative** marks will be awarded for incorrect answers in this section.
17. For each question in **Section 2**, you will be awarded **3 marks** if you darken only the bubble corresponding to the correct answer and **zero mark** if no bubble is darkened. **No negative** marks will be awarded for incorrect answer in this section.

Appropriate way of darkening the bubble for your answer to be evaluated :

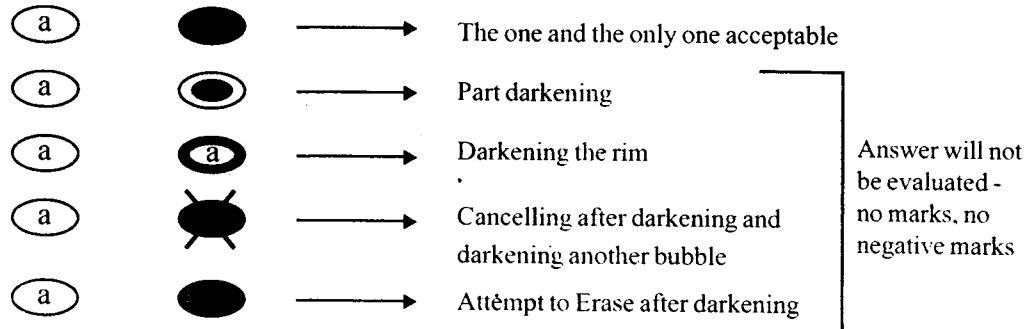


Figure-1 : Correct way of bubbling for valid answer and a few examples of invalid answers. Any other form of partial marking such as ticking or crossing the bubble will be considered invalid.

5	0	4	5	2	3	1
0	0	0	0	0	0	0
1	1	1	1	1	1	1
2	2	2	2	2	2	2
3	3	3	3	3	3	3
4	4	4	4	4	4	4
5	5	5	5	5	5	5
6	6	6	6	6	6	6
7	7	7	7	7	7	7
8	8	8	8	8	8	8
9	9	9	9	9	9	9

Figure-2 : Correct Way of Bubbling your Roll Number on the ORS. (Example Roll Number : 5045231)

Name of the Candidate	Roll Number
KAUSHIK MAJI	5 0 3 6 2 6 8
I have read all instructions and shall abide by them.	I have verified all the information filled by the candidate.
Signature of the Candidate	Signature of the Invigilator